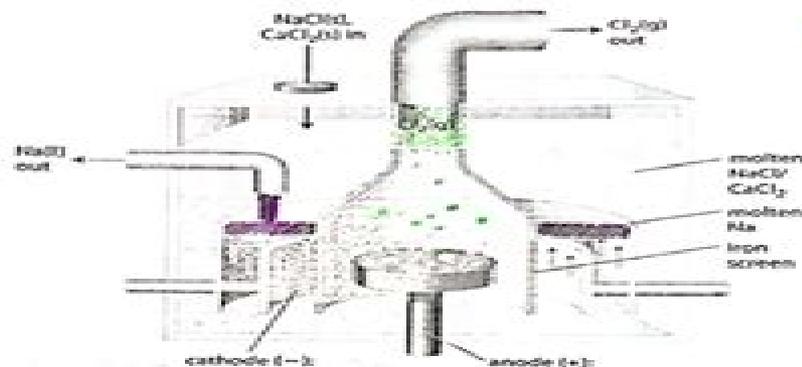


IGCSE – HONORS – ELECTROCHEMISTRY NOTES AND STUDY GUIDE:

Electrolysis of sodium chloride (in the industry)

Industrial electrolysis of molten Sodium Chloride requires extreme temperatures.

- Calcium Chloride is added to the mixture to lower the melting point of Sodium Chloride.
- Molten sodium metal floats in the mixture because of its low density.
- The anode is carbon
- The cathode is iron
- Chlorine gas is collected separately



3.

4. electrolysis of Aluminum oxide (molten)

Electrodes: graphite (carbon)

Electrolyte: molten Alumina (Al₂O₃)

Alumina = Purified Bauxite

Which substance is added to the molten alumina? cryolite

Ions Present: Al³⁺, O²⁻

Half equations:

In the Cathode (Al³⁺ + 3e⁻ → Al) × 4

In the Anode (2O²⁻ → O₂ + 4e⁻) × 3

Overall Equation 4 Al³⁺ + 6 O²⁻ → 4 Al + 3 O₂

Observations:

Cathode: The Molten Aluminum can be observed

Anode: Bubbles of Oxygen gas will appear

Why is it necessary to change the electrodes periodically?

Because the anodes will burn out. Anode reacts with O₂ to form CO₂

Which is the purpose of the cryolite?

To lower the melting point of the Aluminum Oxide



Electrochemistry Study Guide

Camilla Rothe



Electrochemistry Study Guide:

Unveiling the Magic of Words: A Overview of "**Electrochemistry Study Guide**"

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