

Electron Configuration Practice Worksheet **KEY**

In the space below, write the unabbreviated electron configurations of the following elements:

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^4$

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt $[\text{Ar}] 4s^2 3d^7$
- 7) silver $[\text{Kr}] 5s^2 4d^9$
- 8) tellurium $[\text{Kr}] 5s^2 4d^{10} 5p^4$
- 9) radium $[\text{Rn}] 7s^2$
- 10) lawrencium $[\text{Rn}] 7s^2 5f^{14} 6d^1$

Determine what elements are denoted by the following electron configurations:

- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ **Sulfur**
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ **Rubidium**
- 13) $[\text{Kr}] 5s^2 4d^{10} 5p^3$ **Antimony**
- 14) $[\text{Xe}] 6s^2 4f^{14} 5d^6$ **Osmium**
- 15) $[\text{Rn}] 7s^2 5f^{11}$ **Einsteinium**

Determine which of the following electron configurations are not valid:

- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ **should be $3d^{10}$**
- 17) $1s^2 2s^2 2p^6 3s^2 3d^2$ **$3s^2$ is impossible & "p" instead of "d"**
- 18) $[\text{Ra}] 7s^2 5f^8$ **Ra is not a noble gas**
- 19) $[\text{Kr}] 5s^2 4d^{10} 5p^5$ **Valid**
- 20) $[\text{Xe}]$ **Element can't be its own electron configuration.**

Electron Configuration Problems

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Table of Contents Electron Configuration Problems

1. Understanding the eBook Electron Configuration Problems
 - The Rise of Digital Reading Electron Configuration Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Problems
 - Personalized Recommendations
 - Electron Configuration Problems User Reviews and Ratings
 - Electron Configuration Problems and Bestseller Lists

5. Accessing Electron Configuration Problems Free and Paid eBooks
 - Electron Configuration Problems Public Domain eBooks
 - Electron Configuration Problems eBook Subscription Services
 - Electron Configuration Problems Budget-Friendly Options
6. Navigating Electron Configuration Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Problems Compatibility with Devices
 - Electron Configuration Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Problems
 - Highlighting and Note-Taking Electron Configuration Problems
 - Interactive Elements Electron Configuration Problems
8. Staying Engaged with Electron Configuration Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Problems
9. Balancing eBooks and Physical Books Electron Configuration Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Problems
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Electron Configuration Problems
 - Setting Reading Goals Electron Configuration Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configuration Problems
 - Fact-Checking eBook Content of Electron Configuration Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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