

## Electron Configurations Worksheet - Answers

Write the complete ground state electron configurations for the following:

- 1) lithium  $1s^2 2s^1$
- 2) oxygen  $1s^2 2s^2 2p^4$
- 3) calcium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
- 4) titanium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2$
- 5) rubidium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$
- 6) lead  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$
- 7) erbium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{12}$

Write the abbreviated ground state electron configurations for the following:

- 8) helium  $1s^2$  (this one cannot be abbreviated)
- 9) nitrogen  $[\text{He}] 2s^2 2p^3$
- 10) chlorine  $[\text{Ne}] 3s^2 3p^5$
- 11) iron  $[\text{Ar}] 4s^2 3d^6$
- 12) zinc  $[\text{Ar}] 4s^2 3d^{10}$
- 13) barium  $[\text{Xe}] 6s^2$
- 14) polonium  $[\text{Xe}] 6s^2 4f^{14} 5d^{10} 6p^4$

# Electron Configuration Practice With Answers

**Darrell D. Ebbing, Susan M. Schelble**



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