

F325: Equilibria, Energetics and Elements

5.1.3 Acids, Bases & Buffers

117 marks

1. Phenol, C_6H_5OH , is a powerful disinfectant and antiseptic.

Phenol is a weak Brønsted–Lowry acid.



Define the following terms:

- (i) A Brønsted–Lowry acid,

proton donor ✓

(1)

- (ii) A weak acid.

partially dissociates to release H^+ ions
in solution(aq) ✓

(1)

[Total 2 marks]

2. When phenol is mixed with aqueous sodium hydroxide, an acid–base reaction takes place.



In the available spaces,

- label one conjugate acid–base pair as **acid 1** and **base 1**. ✓
- label the other conjugate acid–base pair as **acid 2** and **base 2**. ✓

[Total 1 mark]

3. A solution of phenol in water has a concentration of 4.7 g dm^{-3} .

- (i) Write an expression for the acid dissociation constant, K_a , of phenol.

$$K_a = \frac{[H^+][A^-]}{[HA]} \quad \checkmark$$

(1)

F325 Final Exam Questions

Mike Smith



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