

Chapter 17: Chemical Equilibrium

Worksheet: K<sub>c</sub> and K<sub>p</sub>

1. For the reaction  $2\text{SO}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{SO}_3(g)$

Initial

- 1.  $[\text{SO}_2}] = 0.40 \text{ M}$ ,  $[\text{O}_2}] = 0.20 \text{ M}$ ,  $[\text{SO}_3}] = 0 \text{ M}$
- 2.  $[\text{SO}_2}] = 0.30 \text{ M}$ ,  $[\text{O}_2}] = 0.15 \text{ M}$ ,  $[\text{SO}_3}] = 0.30 \text{ M}$
- 3.  $[\text{SO}_2}] = 0.20 \text{ M}$ ,  $[\text{O}_2}] = 0.10 \text{ M}$ ,  $[\text{SO}_3}] = 0.40 \text{ M}$
- 4.  $[\text{SO}_2}] = 0.10 \text{ M}$ ,  $[\text{O}_2}] = 0.05 \text{ M}$ ,  $[\text{SO}_3}] = 0.50 \text{ M}$
- 5.  $[\text{SO}_2}] = 0.05 \text{ M}$ ,  $[\text{O}_2}] = 0.025 \text{ M}$ ,  $[\text{SO}_3}] = 0.60 \text{ M}$
- 6.  $[\text{SO}_2}] = 0.025 \text{ M}$ ,  $[\text{O}_2}] = 0.0125 \text{ M}$ ,  $[\text{SO}_3}] = 0.70 \text{ M}$
- 7.  $[\text{SO}_2}] = 0.0125 \text{ M}$ ,  $[\text{O}_2}] = 0.00625 \text{ M}$ ,  $[\text{SO}_3}] = 0.80 \text{ M}$
- 8.  $[\text{SO}_2}] = 0.00625 \text{ M}$ ,  $[\text{O}_2}] = 0.003125 \text{ M}$ ,  $[\text{SO}_3}] = 0.90 \text{ M}$
- 9.  $[\text{SO}_2}] = 0.003125 \text{ M}$ ,  $[\text{O}_2}] = 0.0015625 \text{ M}$ ,  $[\text{SO}_3}] = 0.975 \text{ M}$
- 10.  $[\text{SO}_2}] = 0.0015625 \text{ M}$ ,  $[\text{O}_2}] = 0.00078125 \text{ M}$ ,  $[\text{SO}_3}] = 0.9875 \text{ M}$

Answers

- 11.  $K_c = 1.6$
- 12.  $K_c = 1.6$
- 13.  $K_c = 1.6$
- 14.  $K_c = 1.6$
- 15.  $K_c = 1.6$
- 16.  $K_c = 1.6$
- 17.  $K_c = 1.6$
- 18.  $K_c = 1.6$
- 19.  $K_c = 1.6$
- 20.  $K_c = 1.6$
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- 97.  $K_c = 1.6$
- 98.  $K_c = 1.6$
- 99.  $K_c = 1.6$
- 100.  $K_c = 1.6$

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**Lauren Gardner**



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