

Reaction Rates and Chemical Equilibrium

A. Factors That Affect the Rate of a Reaction

A.1 Effect of Temperature

Test Tube Temperature (3)

Observations (4)

1. 10 °C

Slow fizzing, clear solution after 2 min 38 sec

2. 58 °C

vigorous fizzing/bubbles, clear solution in

Which test tube cleared first (4)?

48 sec

Q1 How did an increase in temperature affect the rate of a reaction?

A.2 Effect of Changing Reactant Concentration

Test Tube	Concentration of HCl	Initial time (2)	Final time (3)	Total time (3) - (2)
1.	1.0 M HCl	0	17 min 5 sec	
2.	2.0 M HCl	0	4 min 38 sec	
3.	3.0 M HCl	0	1 min 20 sec	

Rank the test tubes in order of slowest to fastest. _____ < _____ < _____

Q2 How did an increase in the concentration of HCl cause a difference in the rates of reaction of HCl with Mg?

Lab Report On Reaction Rates And Equilibrium

BM King



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Reaction Rates for High-temperature Air with Carbon and Sodium Impurities Mina L. Carnicom, 1968 The values used by a number of investigators for the rate constants of high temperature greater than or equal to 1000 K homogeneous gaseous reactions involving species of the elements nitrogen oxygen carbon and sodium have been compiled and are presented in tabular form Included are reactions involving neutral species charged species free electrons some species in excited electronic or vibrational states and radiative processes *Applied Mechanics Reviews*, 1961 EPA-625/6, 1990-07

Assessing the Geochemical Fate of Deep-well-injected Hazardous Wastes, 1990 **Uranium Production Technology** Charles D. Harrington, Archie E. Ruehle, 1959 Research Report - Avco Everett Research Laboratory Avco Corporation. Everett Research Laboratory, 1969 *Laboratory manual* William McPherson, 1913 **A Course of Laboratory Experiments on Physico-chemical Principles** Miles Standish Sherrill, 1923 **Fossil Energy Update**, 1986

Energy Research Abstracts, 1977 Semiannual with semiannual and annual indexes References to all scientific and technical literature coming from DOE its laboratories energy centers and contractors Includes all works deriving from DOE other related government sponsored information and foreign nonnuclear information Arranged under 39 categories e g Biomedical sciences basic studies Biomedical sciences applied studies Health and safety and Fusion energy Entry gives bibliographical information and abstract Corporate author subject report number indexes **Journal of Gas Lighting**, 1924 Gas Journal, 1924 *Annual Report of the Director of the Geophysical Laboratory* Carnegie Institution of Washington. Geophysical Laboratory, 1930 **Transfer Processes in Cohesive Sediment Systems** W. R. Parker, 2012-12-06 In the transfer of chemical species through the aquatic environment the association with fine sediment particles is often of crucial importance The nature permanence and kinetics of this association is in many cases unknown yet it is central to any effort concerned with predicting the pathways of pollutant transfer and the fluxes along them It is often unclear to precisely what surface if any a pollutant is attached The transfer of species between surfaces and solution may be chemically or biochemically controlled These processes may take place within a host framework which has a time and space dependent structure The natural straining of this framework resulting from physical and chemical processes moves porewater and the dissolved species in it Thus to adequately predict the natural transfer of species within a cohesive sediment system the physical biochemical and chemical processes must be dynamically coupled This informal colloquium examined topics relevant to the association of pollutant and sediment These included the mineral particle surfaces and the surrounding ionic associations methods of describing the particles and the physical chemical biochemical and biological processes operating in association with fine sediment substrates The physical processes of sediment transport were deliberately excluded **Reaction Rate Compilation for the H-O-N System** Gilbert S. Bahn, 1968 **Journal of the American Chemical Society** American Chemical Society, 1928 Proceedings of the Society are included in v 1 59 1879 1937

Nuclear Science Abstracts ,1957 *Laboratory Exercises in Inorganic Chemistry* James Flack Norris, Kenneth
Lamartine Mark,1922 Journal of Gas Lighting and Water Supply ,1924 **Catalog Issue** Pratt Institute. School of
Science and Technology,1926

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