

Part A: Concentration of Stock is  $1.94 \times 10^{-4}$  M  
 Table 1. Preparation of Standard Solutions

Standard	Volume of Stock Solution (mL)	Final Volume of Diluted Solution (mL)	Concentration of Diluted Solutions (M)
1		50.00	$1.94 \times 10^{-4}$
2		50.00	$3.87 \times 10^{-4}$
3		50.00	$9.70 \times 10^{-4}$
4		50.00	$1.36 \times 10^{-3}$

Show your calculation for at least one value of volume of stock solution.

Part B:

Record the wavelength that had the highest absorbance for this dye: 525 nm

Part C:

Table 2. Absorbance data for standards and unknown.

Trial	Concentration (mol/L)	Absorbance (no unit)
1	$1.94 \times 10^{-4}$	0.032
2	$3.87 \times 10^{-4}$	0.084
3	$9.70 \times 10^{-4}$	0.221
4	$1.36 \times 10^{-3}$	0.268
Unknown Letter <u>          </u>	(Answer from Q. #3, next page)	0.156
Equation of line:		

Analysis

Answer question 1 while working on the procedure for Part C (step 3):

1. Estimate the concentration of your unknown by comparing it to your standard solutions: (Circle one)

- |                                       |                                      |
|---------------------------------------|--------------------------------------|
| a. Less concentrated than standard #1 | d. Between standard #3 and #4        |
| b. Between standard #1 and #2         | e. More concentrated than standard # |
| c. Between standard #2 and #3         |                                      |

# Laboratory Investigation 30a Answers

**M Lipman**



**Laboratory Investigation 30a Answers:**

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