

Stoichiometry - Mixed Conversions Answer Key

1. Use the equation below to answer the following questions:



a. How many grams of NH_3 are made from 31L of N_2 gas?

$$\frac{31\text{L N}_2}{1} \times \frac{1\text{mole N}_2}{22.4\text{L N}_2} \times \frac{2\text{mole NH}_3}{1\text{mole N}_2} \times \frac{17\text{g NH}_3}{1\text{mole NH}_3} = 47.3\text{g NH}_3$$

b. How many molecules of H_2 are needed to produce 60g of NH_3 ?

$$\frac{60\text{g NH}_3}{1} \times \frac{1\text{mole NH}_3}{17\text{g NH}_3} \times \frac{1\text{mole N}_2}{2\text{mole NH}_3} \times \frac{3\text{mole H}_2 \times 6.02 \times 10^{23}\text{ molecules H}_2}{1\text{mole H}_2} = 1.22 \times 10^{24}\text{ molecules H}_2$$

c. How many moles of N_2 are needed to completely react with 87L of H_2 gas?

$$\frac{87\text{L}}{1} \times \frac{1\text{mole H}_2}{22.4\text{L H}_2} \times \frac{1\text{mole N}_2}{3\text{mole H}_2} = 1.27\text{mole N}_2$$

2. Butane reacts with oxygen to form water and carbon dioxide.



a. How many molecules of H_2O are produced with 20.7L of O_2 gas that are completely used during the combustion of butane?

$$\frac{20.7\text{L O}_2}{1} \times \frac{1\text{mole O}_2}{22.4\text{L O}_2} \times \frac{10\text{mole H}_2\text{O}}{13\text{mole O}_2} \times \frac{6.02 \times 10^{23}\text{ molecules H}_2\text{O}}{1\text{mole H}_2\text{O}} = 4.17 \times 10^{23}\text{ molecules H}_2\text{O}$$

b. If 48g of H_2O are collected from the burning of butane, then how many moles of butane were reacted with oxygen?

$$\frac{48\text{g H}_2\text{O}}{1} \times \frac{1\text{mole H}_2\text{O}}{18\text{g H}_2\text{O}} \times \frac{2\text{mole C}_4\text{H}_{10}}{10\text{mole H}_2\text{O}} = 0.53\text{mole C}_4\text{H}_{10}$$

3. Potassium reacts with water to produce potassium hydroxide and hydrogen gas.



a. How many grams of H_2 will be produced if 7.22×10^{23} atoms of potassium react completely with water?

$$\frac{7.22 \times 10^{23}\text{ atoms K}}{1} \times \frac{1\text{mole K}}{6.02 \times 10^{23}\text{ atoms K}} \times \frac{1\text{mole H}_2}{2\text{mole K}} \times \frac{2\text{g H}_2}{1\text{mole H}_2} = 0.60\text{g H}_2$$

b. How many atoms of K are needed to completely react with 3.22×10^{23} molecules of water?

$$\frac{3.22 \times 10^{23}\text{ molecules H}_2\text{O}}{1} \times \frac{1\text{mole H}_2\text{O}}{6.02 \times 10^{23}\text{ molecules H}_2\text{O}} \times \frac{2\text{mole K}}{2\text{mole H}_2\text{O}} \times \frac{6.02 \times 10^{23}\text{ atoms K}}{1\text{mole K}} = 3.22 \times 10^{23}\text{ atoms K}$$

Mixed Stoichiometry Worksheet

Raffaella Di Napoli



Mixed Stoichiometry Worksheet:

The Enigmatic Realm of **Mixed Stoichiometry Worksheet**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing in short supply of extraordinary. Within the captivating pages of **Mixed Stoichiometry Worksheet** a literary masterpiece penned with a renowned author, readers set about a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting affect the hearts and minds of those that partake in its reading experience.

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