

Unit 7: Balancing Chemical Reactions- Worksheet 2

Balance the following equations by inserting the proper coefficients.

- $2 \text{SO}_2 + 1 \text{O}_2 \rightarrow 2 \text{SO}_3$ **Synthesis**
 - $2 \text{CH}_4 + 3 \text{O}_2 \rightarrow 2 \text{CO} + 4 \text{H}_2\text{O}$ **Single R**
 - $2 \text{P} + 3 \text{Cl}_2 \rightarrow 2 \text{PCl}_3$ **Synthesis**
 - $2 \text{CO} + 1 \text{O}_2 \rightarrow 2 \text{CO}_2$ **Synthesis**
 - $2 \text{CH}_4 + 1 \text{O}_2 \rightarrow 2 \text{CH}_3\text{OH}$ **Synthesis**
 - $2 \text{Li} + 1 \text{Br}_2 \rightarrow 2 \text{LiBr}$ **Synthesis**
 - $2 \text{Al}_2\text{O}_3 \rightarrow 4 \text{Al} + 3 \text{O}_2$ **Decomp**
 - $2 \text{Na} + 2 \text{H}_2\text{O} \rightarrow 2 \text{NaOH} + 1 \text{H}_2$ **Single R**
 - $6 \text{CO}_2 + 6 \text{H}_2\text{O} \rightarrow 1 \text{C}_6\text{H}_{12}\text{O}_6 + 6 \text{O}_2$ **Rev. Comb**
 - $1 \text{H}_2\text{SO}_4 + 2 \text{NaCl} \rightarrow 2 \text{HCl} + 1 \text{Na}_2\text{SO}_4$ **Double R**
- Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:
- $3 \text{H}_2 + 1 \text{SO}_2 \rightarrow 1 \text{H}_2\text{S} + 2 \text{H}_2\text{O}$ **Double R**
 - $2 \text{CaCO}_3 + 2 \text{SO}_2 + 1 \text{O}_2 \rightarrow 2 \text{CaSO}_4 + 2 \text{CO}_2$ **No**
 - $2 \text{AgNO}_3 + 1 \text{CaCl}_2 \rightarrow 2 \text{AgCl} + 1 \text{Ca}(\text{NO}_3)_2$ **Double R**
 - $2 \text{HCl} + 1 \text{Ba}(\text{OH})_2 \rightarrow 1 \text{BaCl}_2 + 2 \text{H}_2\text{O}$ **Double R**
 - $1 \text{H}_3\text{PO}_4 + 3 \text{NaOH} \rightarrow 1 \text{Na}_3\text{PO}_4 + 3 \text{H}_2\text{O}$ **Double R**
 - $1 \text{Pb}(\text{NO}_3)_2 + 2 \text{KI} \rightarrow 1 \text{PbI}_2 + 2 \text{KNO}_3$ **Double R**
 - $3 \text{CuO} + 2 \text{NH}_3 \rightarrow 1 \text{N}_2 + 3 \text{Cu} + 3 \text{H}_2\text{O}$ **Single R**
 - $1 \text{C}_2\text{H}_5\text{OH} + 3 \text{O}_2 \rightarrow 2 \text{CO}_2 + 3 \text{H}_2\text{O}$ **Com**
 - $2 \text{C}_2\text{H}_6 + 3 \text{O}_2 \rightarrow 2 \text{CH}_3\text{COOH} + 2 \text{H}_2\text{O}$ **No**
 - $3 \text{NO}_2 + 1 \text{H}_2\text{O} \rightarrow 2 \text{HNO}_3 + 1 \text{NO}$ **No**

Modeling Chemistry U7 Ws V2 Key

Frank Lees



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Table of Contents Modeling Chemistry U7 Ws V2 Key

1. Understanding the eBook Modeling Chemistry U7 Ws V2 Key
 - The Rise of Digital Reading Modeling Chemistry U7 Ws V2 Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U7 Ws V2 Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry U7 Ws V2 Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U7 Ws V2 Key
 - Personalized Recommendations
 - Modeling Chemistry U7 Ws V2 Key User Reviews and Ratings
 - Modeling Chemistry U7 Ws V2 Key and Bestseller Lists

5. Accessing Modeling Chemistry U7 Ws V2 Key Free and Paid eBooks
 - Modeling Chemistry U7 Ws V2 Key Public Domain eBooks
 - Modeling Chemistry U7 Ws V2 Key eBook Subscription Services
 - Modeling Chemistry U7 Ws V2 Key Budget-Friendly Options
6. Navigating Modeling Chemistry U7 Ws V2 Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry U7 Ws V2 Key Compatibility with Devices
 - Modeling Chemistry U7 Ws V2 Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry U7 Ws V2 Key
 - Highlighting and Note-Taking Modeling Chemistry U7 Ws V2 Key
 - Interactive Elements Modeling Chemistry U7 Ws V2 Key
8. Staying Engaged with Modeling Chemistry U7 Ws V2 Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry U7 Ws V2 Key
9. Balancing eBooks and Physical Books Modeling Chemistry U7 Ws V2 Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry U7 Ws V2 Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry U7 Ws V2 Key
 - Setting Reading Goals Modeling Chemistry U7 Ws V2 Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry U7 Ws V2 Key
 - Fact-Checking eBook Content of Modeling Chemistry U7 Ws V2 Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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