

$$\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100\%$$

↳ from B-C-A chart

## Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation.
- Identify the given (with units) and what you want to find (with units).
- Show set up with units. Give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



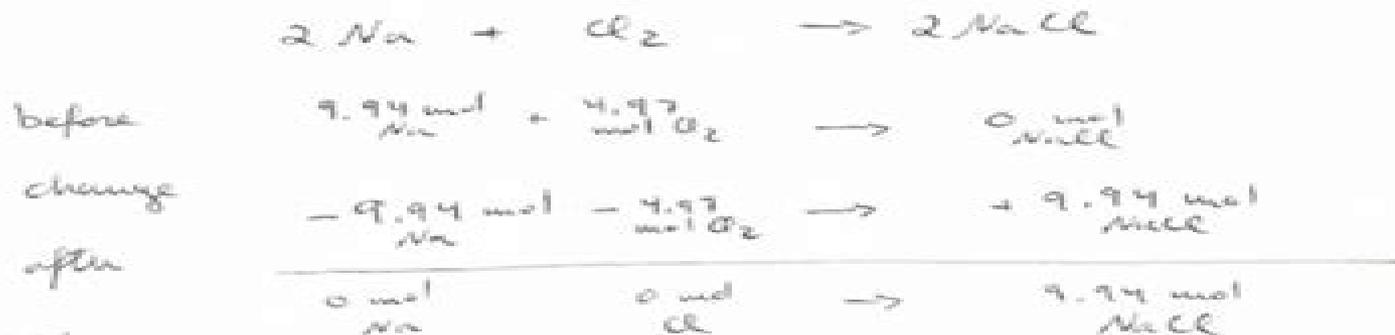
1 mol H<sub>2</sub>O : 18.02 g/mol

36 g  $\times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g}} = 1.99 \text{ mol H}_2\text{O}$

1 mol H<sub>2</sub> : 2.02 g

2 mol H<sub>2</sub> = 4.04 g H<sub>2</sub>

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



1 mol NaCl = 58.44 g  
 $581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.44 \text{ g}} = 9.94 \text{ mol NaCl}$

1 mol Na = 22.99 g

9.94 mol Na  $\times$  22.99 g = 228.52 g Na

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## **Table of Contents Modeling Chemistry U8 Ws V2 Key**

1. Understanding the eBook Modeling Chemistry U8 Ws V2 Key
  - The Rise of Digital Reading Modeling Chemistry U8 Ws V2 Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U8 Ws V2 Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry U8 Ws V2 Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U8 Ws V2 Key
  - Personalized Recommendations

- Modeling Chemistry U8 Ws V2 Key User Reviews and Ratings
- Modeling Chemistry U8 Ws V2 Key and Bestseller Lists
- 5. Accessing Modeling Chemistry U8 Ws V2 Key Free and Paid eBooks
  - Modeling Chemistry U8 Ws V2 Key Public Domain eBooks
  - Modeling Chemistry U8 Ws V2 Key eBook Subscription Services
  - Modeling Chemistry U8 Ws V2 Key Budget-Friendly Options
- 6. Navigating Modeling Chemistry U8 Ws V2 Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Modeling Chemistry U8 Ws V2 Key Compatibility with Devices
  - Modeling Chemistry U8 Ws V2 Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry U8 Ws V2 Key
  - Highlighting and Note-Taking Modeling Chemistry U8 Ws V2 Key
  - Interactive Elements Modeling Chemistry U8 Ws V2 Key
- 8. Staying Engaged with Modeling Chemistry U8 Ws V2 Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modeling Chemistry U8 Ws V2 Key
- 9. Balancing eBooks and Physical Books Modeling Chemistry U8 Ws V2 Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modeling Chemistry U8 Ws V2 Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry U8 Ws V2 Key
  - Setting Reading Goals Modeling Chemistry U8 Ws V2 Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry U8 Ws V2 Key
  - Fact-Checking eBook Content of Modeling Chemistry U8 Ws V2 Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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