

Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H₂O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	58.5 g

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	84.0

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	159.6

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D Keegan



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1. Understanding the eBook Modeling Chemistry Unit 11 Answer Key
 - The Rise of Digital Reading Modeling Chemistry Unit 11 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 11 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 11 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 11 Answer Key

- Personalized Recommendations
 - Modeling Chemistry Unit 11 Answer Key User Reviews and Ratings
 - Modeling Chemistry Unit 11 Answer Key and Bestseller Lists
5. Accessing Modeling Chemistry Unit 11 Answer Key Free and Paid eBooks
- Modeling Chemistry Unit 11 Answer Key Public Domain eBooks
 - Modeling Chemistry Unit 11 Answer Key eBook Subscription Services
 - Modeling Chemistry Unit 11 Answer Key Budget-Friendly Options
6. Navigating Modeling Chemistry Unit 11 Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - Modeling Chemistry Unit 11 Answer Key Compatibility with Devices
 - Modeling Chemistry Unit 11 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 11 Answer Key
 - Highlighting and Note-Taking Modeling Chemistry Unit 11 Answer Key
 - Interactive Elements Modeling Chemistry Unit 11 Answer Key
8. Staying Engaged with Modeling Chemistry Unit 11 Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 11 Answer Key
9. Balancing eBooks and Physical Books Modeling Chemistry Unit 11 Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 11 Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry Unit 11 Answer Key
- Setting Reading Goals Modeling Chemistry Unit 11 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry Unit 11 Answer Key

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- Fact-Checking eBook Content of Modeling Chemistry Unit 11 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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