

MODERN BIOLOGY

Ch. 2:1 Chemistry

MATTER: Chemical changes in matter are essential to all life processes.

- **Matter:** Anything that occupies space and has mass.
- **Mass:** The quantity of matter an object has.

ELEMENTS AND ATOMS

- Elements are substances that cannot be broken down chemically into simpler kinds of matter.
- >90% of all kinds of living things are made of combinations of 4 elements:
 - OXYGEN (O₂)
 - CARBON (C)
 - HYDROGEN (H₂)
 - NITROGEN (N)
- Periodic Table of the Elements
 - Atomic number:
 - Chemical symbol:
 - Atomic mass:
- **Atom:** the simplest particle of an element that retains all of the properties of that element
- Protons:
- Neutrons:
- Electrons:
- Orbital:
- Ion:
- isotope:

COMPOUNDS

- The physical and chemical properties differ between the compounds and elements that compose them.
For example:
 - Sodium (Na) is a soft metal that explodes when placed in water.
 - Chlorine (Cl) is a green poisonous gas.
 - Sodium Chloride (NaCl) is table salt.
- All atoms are neutral.
- Most atoms are unstable. (Noble gases are stable as atoms. They do not react.)
- Atoms combine (react) to become stable.
- Atoms become stable when their outermost energy level is "full" of electrons. They will:
 - share electrons equally
 - share electrons unequally
 - transfer electrons completely with another atom (gain or loss but no sharing)
- **The number and arrangement of their electrons determine the way that elements combine.**
- Elements are arranged on the periodic table by the number of valence electrons and the number of energy levels that they have. Understand the periodic table so you can learn how compounds are made.
- **Covalent Bonds:** electrons are either shared equally or unequally (co-sharing; -valent; outermost electrons). Atoms in a molecule stay together to maintain stability.
- **Ionic Bonds:** electron(s), given up or accepted completely; elements in the compound ionize (ionic

bond). Atoms (ions) in a compound stay together because opposite charges attract.

Review Questions:

1. Carbon has an atomic number of 6. How many p⁺ and e⁻ does an atom of carbon have?
2. Sodium is in family I, therefore, it has 1 valence e⁻. What will Na⁺ do to become stable?

Challenge

3. If Aluminum gives up it's 3 valence e⁻ to chlorine atoms, what are the resulting charges on the ions of Al and each Cl (AlCl₃)?
4. Oxygen is in family VI. It has 6 valence e⁻ and needs 2 more to become stable. Draw an electron configuration diagram (dot diagram) and count out the e⁻ that make each atom stable. (#3, pg. 10 of study guide)
5. An acetylene torch gives off lots of energy. In fact, it is used to cut steel plate. The formula for acetylene is C₂H₂, and there is a triple bond between the carbon atoms. Draw in the e⁻ where the ? are located below. (C = Family IV; H = Family I)



MODERN BIOLOGY

Ch. 2:2 ENERGY

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Modern Biology Ch 49 Study Guide Answers

LP Steffe



Modern Biology Ch 49 Study Guide Answers:

Modern Biology Ch 49 Study Guide Answers Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Modern Biology Ch 49 Study Guide Answers**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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