

## Chemistry – Unit 7 Worksheet 3

Write balanced chemical equations for the following reactions.

Note: All coefficients of 1 are omitted; in the first 10 items, the coefficients are all 1.

1. Ammonia ( $\text{NH}_3$ ) reacts with hydrogen chloride to form ammonium chloride.  
 $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$

2. Calcium carbonate decomposes upon heating to form calcium oxide and carbon dioxide.  
 $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$

3. Barium oxide reacts with water to form barium hydroxide.  
 $\text{BaO} + \text{H}_2\text{O} \rightarrow \text{Ba}(\text{OH})_2$

4. Acetaldehyde ( $\text{CH}_3\text{CHO}$ ) decomposes to form methane ( $\text{CH}_4$ ) and carbon monoxide.  
 $\text{CH}_3\text{CHO} \rightarrow \text{CH}_4 + \text{CO}$

5. Zinc reacts with copper (II) nitrate to form zinc nitrate and copper.  
 $\text{Zn} + \text{Cu}(\text{NO}_3)_2 \rightarrow \text{Zn}(\text{NO}_3)_2 + \text{Cu}$

6. Calcium sulfite decomposes when heated to form calcium oxide and sulfur dioxide.  
 $\text{CaSO}_3 \rightarrow \text{CaO} + \text{SO}_2$

7. Iron reacts with sulfuric acid ( $\text{H}_2\text{SO}_4$ ) to form iron (II) sulfate and hydrogen gas.  
 $\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$

8. A nitrogen containing carbon compound,  $\text{C}_2\text{H}_6\text{N}_2$ , decomposes to form ethane,  $\text{C}_2\text{H}_6$ , and nitrogen gas.  
 $\text{C}_2\text{H}_6\text{N}_2 \rightarrow \text{C}_2\text{H}_6 + \text{N}_2$

9. Phosgene,  $\text{COCl}_2$ , is formed when carbon monoxide reacts with chlorine gas.  
 $\text{CO} + \text{Cl}_2 \rightarrow \text{COCl}_2$

# Modeling Chemistry Unit 7 3

**Ying Liu**



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web mar 22 2023 new class 1c boiler is a test about this specific boiler it sounds an easy one but it all depends on how good you are with such appliances take it up and see