

Kinetics: "Iodine Clock" Lab Report

Introduction: In order to describe the chemical kinetics of a reaction, it is desirable to determine how the rate of reaction varies as the reaction progresses. The rate law is a mathematical equation that describes the progress of the reaction and has the following general form for the reaction $aA + bB \rightarrow cC + dD$:

$$\text{rate} = -d[A]/(a(dt)) = k[A]^m[B]^n$$

[A] and [B] are concentrations of the reactants and k is the rate constant. The variables m, n and k can only be determined experimentally. In this experiment, the oxidation reaction of iodide by peroxydisulfate is performed. The rate law for this reaction is as follows:

$$\text{rate} = -\Delta[S_2O_8^{2-}]/\Delta t = k[I^-]^m[S_2O_8^{2-}]^n$$

This lab provides an opportunity to understand different concepts of chemical kinetics such as the reaction rate, rate constant, and reaction order. In this lab-- using several mixtures of the iodide and peroxydisulfate solutions-- it is possible to calculate the reaction order and the reaction constant of the chemical reaction. By manipulation of temperature, it is also possible to understand the temperature dependence of the rate constant as well as determine the activation energy of the reaction using an Arrhenius plot. Understanding chemical kinetics is important because of its practical applications. For example, understanding rates of reaction is important in the manufacturing of ammonia for fertilizer using the Haber process. If the reaction to produce ammonia is conducted at a low temperature, the rate of reaction will be too slow. Because the manufacturers want to produce as much fertilizer as possible, they must manipulate the temperature to produce a satisfactory rate.

Experimental: For the first portion of this experiment, 5 "Mixtures" are prepared, with varying concentrations of solutions; the concentrations are summarized in the following table:

Mixture #	I ⁻ (0.20M)	SO ₄ ²⁻ (0.33M)	S ₂ O ₈ ²⁻ (0.01M)	Starch (drops)	H ₂ O	S ₂ O ₈ ²⁻ (0.14M)*
1	50	25	10	10	5	10
2	25	30	10	10	25	10
3	10	33	10	10	37	10
4	10	27	10	10	28	25
5	10	16	10	10	14	50

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