

Molar Mass Practice 1.8 – Answer Key

Objectives: - Calculate molar mass using the periodic table

Directions: Find the molar mass of the following atoms, molecules, or formula units. Round masses to the hundredths place. Don't forget your units! You must show your work!

1. Cl_2

$$2 \times 35.45 \text{ g/mol} = \boxed{70.90 \text{ g/mol}}$$

9. $(\text{NH}_4)_2\text{S}$

$$2 \times 14.01 + 8 \times 1.01 + 32.07 = \boxed{98.17 \text{ g/mol}}$$

2. Fe

$$55.85 \text{ g/mol}$$

10. $\text{C}_2\text{H}_6\text{O}$

$$\text{C: } 2 \times 12.01 = 24.02$$

$$\text{H: } 6 \times 1.01 = 6.06$$

$$\text{O: } 1 \times 16.00 = 16.00$$

$$\boxed{46.08 \text{ g/mol}}$$

3. CO_2

$$\text{C: } 1 \times 12.01 = 12.01$$

$$\text{O: } 2 \times 16.00 = 32.00$$

$$\boxed{44.01 \text{ g/mol}}$$

11. $3\text{Mg}(\text{NO}_3)_2$

$$3\text{Mg: } 3 \times 24.31 = 72.93$$

$$\text{N: } 6 \times 14.01 = 84.06$$

$$\text{O: } 18 \times 16.00 = 288.00$$

$$\boxed{445.00 \text{ g/mol}}$$

4. LiCl

$$\text{Li: } 1 \times 6.94 = 6.94$$

$$\text{Cl: } 1 \times 35.45 = 35.45$$

$$\boxed{42.39 \text{ g/mol}}$$

12. H_2SO_4

$$2 \times 1.01 + 32.07 + (4 \times 16.00) = \boxed{98.09 \text{ g/mol}}$$

5. O_2

$$\text{O: } 2 \times 16.00 = 32.00$$

$$\boxed{32.00 \text{ g/mol}}$$

13. $\text{Mg}_3(\text{PO}_4)_2$

$$3\text{Mg: } 3 \times 24.31 = 72.93$$

$$\text{P: } 2 \times 30.97 = 61.94$$

$$\text{O: } 8 \times 16.00 = 128.00$$

$$\boxed{262.87 \text{ g/mol}}$$

6. Iron(III) oxide



$$2 \times 55.85 + 3 \times 16.00 = \boxed{159.7 \text{ g/mol}}$$

14. Dihydrogen monoxide



$$2 \times 1.01 + 16.00 = \boxed{18.02 \text{ g/mol}}$$

7. Hg_2Br_2

$$\boxed{536.00 \text{ g/mol}}$$

15. $\text{Zn}(\text{NO}_3)_2$

$$\boxed{236.4 \text{ g/mol}}$$

8. Lead(II) hydroxide



$$\boxed{241.2 \text{ g/mol}}$$

16. Aluminum sulfate



$$\boxed{698 \text{ g/mol}}$$

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Brendan G. Carr



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