

General Chemistry Homework Answer Key

Textbook: OpenStax Chemistry: Atoms First 2e

Chapter 3: Electronic Structure and Periodic Properties of Elements

TA Reference Answers with Step-by-Step Explanations

3.1 Electromagnetic Energy

- 2 – Define frequency, wavelength, and amplitude.

Answer:

Frequency (ν) is the number of wave cycles that pass a point per second, measured in hertz (Hz).

Wavelength (λ) is the distance between successive crests of a wave, measured in meters (m).

Amplitude is the height of the wave from the center line to the crest or trough, indicating intensity.

Stepwise Procedure:

1. Define each term conceptually.
2. Include appropriate units.
3. Clarify differences among the properties.

- 4 – What is the relationship between frequency and energy?

Answer:

The energy (E) of electromagnetic radiation is directly proportional to its frequency (ν): $E = h\nu$, where h is Planck's constant (6.626×10^{-34} J·s).

Stepwise Procedure:

1. State the energy-frequency equation.
2. Define h and ν .
3. Explain that higher frequency means higher energy.

3.2 The Bohr Model

- 8 – What are the key features of the Bohr model of the hydrogen atom?

Answer:

1. Electrons orbit the nucleus in fixed paths (energy levels).
2. Only certain orbits are allowed with quantized energies.
3. Electrons absorb or emit energy when they move between levels.

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