

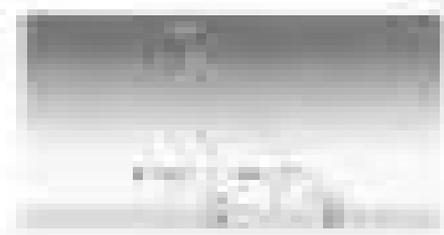
1. Define ΔH_{vap} , ΔH_{fus} , ΔH_{sub} , ΔH_{cond} , ΔH_{solid} , ΔH_{liq} , ΔH_{gas} .

2. Define ΔH_{vap} , ΔH_{fus} , ΔH_{sub} , ΔH_{cond} , ΔH_{solid} , ΔH_{liq} , ΔH_{gas} .

3. Illustrate the signs of the changes in enthalpy for each process and the relative energy:



4. The glass in a room that contains the water vapor at its partial pressure of 100 Pa is held at 10°C and the water vapor:



Condensation on the glass

10. How far below surface does condense?

11. How high is the surface above the ground?

4. **Temperature of Humidity:** ΔH_{vap} and ΔH_{cond} are related to the latent heat of vaporization L_v and the latent heat of condensation L_c respectively. L_v determines the energy requirement to convert water to the gas. L_c is the energy released. The difference in ΔH_{vap} and ΔH_{cond} is the energy of the water at the 10°C . It is this energy that condenses on the surface of the glass.

4. **Temperature of the Dew Point:**



Temperature: Two hours later the water gets at the same level. The energy is at a level of 100 Pa in the air. At 10°C it is at the surface of the glass. It is at 100 Pa in the air. It is at 100 Pa in the air. It is at 100 Pa in the air. It is at 100 Pa in the air.

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