

1. D

**Explanation:** The orbital that must be filled last is the one with the highest  $(n + l)$  value. For  $6s$  orbital,  $(n + l) = 6$ ; for  $5d$ ,  $(n + l) = 7$ ; for  $4f$ ,  $(n + l) = 7$ ; and for  $6p$ ,  $(n + l) = 7$ . Since options (b), (c), and (d) all have the same  $(n + l)$  value, we apply the rule that the one with the lower  $n$  value will be filled first. Going with these, it can be said that  $6p$  orbital will be filled last.

2. C

**Explanation:** Based on the  $(n + l)$  values calculated in No. 1,  $6s$  is the 1st priority to be filled with electrons, followed by  $4f$ , then  $5d$ , and lastly,  $6p$ .

3. A

**Explanation:** The set of quantum numbers  $(1, 1, 0, +\frac{1}{2})$  is impossible because an  $l$  value equal to 1 suggests that the orbital is a  $p$  orbital. Unfortunately, there is no  $p$  orbital in  $n = 1$  (that's why there is no  $1p^1$ ), hence impossible!

4. C

**Explanation:** Based on the mnemonic device that we are using, the correct order of orbitals to use corresponds to that in option (c). Hence, it is the correct answer.

5. A

**Explanation:** The only  $d$  orbital in the electron configuration of Mn is the  $3d$  orbital. Therefore,  $n = 3$ . Furthermore, there is only one  $l$  value for  $d$  orbital, and that is 2. If we assign  $m_l$  values to the five subshells of  $3d$  from  $-2$  to  $+2$ , the labeled electron in the figure will fall under  $-1$ . Therefore, the correct answer is option (a).

# History Of Atoms Answer Key

**Hussin A.Rothana**



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