

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the to systems, LeChatelier's principle and equilibrium. Through applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
2. Write a balanced equation for the reaction in part 2.
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
3. Write a balanced equation for the reaction in part 3.
 $\text{Co}^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
4. In part 1, what is the conjugate acid and base? Write the K_{eq} for the reaction. Is the reaction exothermic or endothermic? Explain the effect of temperature on the reaction. How does the color of the solution change as a function of temperature? Explain the color change.
5. What is the conjugate acid and base? Write the K_{eq} for the reaction. Is the reaction exothermic or endothermic? Explain the effect of temperature on the reaction. How does the color of the solution change as a function of temperature? Explain the color change.
6. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
 $\text{Fe}^{3+}(\text{aq}) + 3\text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$
7. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
 $\text{Fe}^{3+}(\text{aq}) + 3\text{OH}^{-}(\text{aq}) \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$
8. Write the equation for the reaction between Fe^{3+} and OH^{-} ions.
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Investigating Chemical Equilibrium Lab 12a

David Kirk



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Table of Contents Investigating Chemical Equilibrium Lab 12a

1. Understanding the eBook Investigating Chemical Equilibrium Lab 12a
 - The Rise of Digital Reading Investigating Chemical Equilibrium Lab 12a
 - Advantages of eBooks Over Traditional Books
2. Identifying Investigating Chemical Equilibrium Lab 12a
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Investigating Chemical Equilibrium Lab 12a
 - User-Friendly Interface
4. Exploring eBook Recommendations from Investigating Chemical Equilibrium Lab 12a
 - Personalized Recommendations
 - Investigating Chemical Equilibrium Lab 12a User Reviews and Ratings
 - Investigating Chemical Equilibrium Lab 12a and Bestseller Lists
5. Accessing Investigating Chemical Equilibrium Lab 12a Free and Paid eBooks

- Investigating Chemical Equilibrium Lab 12a Public Domain eBooks
 - Investigating Chemical Equilibrium Lab 12a eBook Subscription Services
 - Investigating Chemical Equilibrium Lab 12a Budget-Friendly Options
6. Navigating Investigating Chemical Equilibrium Lab 12a eBook Formats
- ePub, PDF, MOBI, and More
 - Investigating Chemical Equilibrium Lab 12a Compatibility with Devices
 - Investigating Chemical Equilibrium Lab 12a Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Investigating Chemical Equilibrium Lab 12a
 - Highlighting and Note-Taking Investigating Chemical Equilibrium Lab 12a
 - Interactive Elements Investigating Chemical Equilibrium Lab 12a
8. Staying Engaged with Investigating Chemical Equilibrium Lab 12a
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Investigating Chemical Equilibrium Lab 12a
9. Balancing eBooks and Physical Books Investigating Chemical Equilibrium Lab 12a
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Investigating Chemical Equilibrium Lab 12a
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Investigating Chemical Equilibrium Lab 12a
- Setting Reading Goals Investigating Chemical Equilibrium Lab 12a
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Investigating Chemical Equilibrium Lab 12a
- Fact-Checking eBook Content of Investigating Chemical Equilibrium Lab 12a
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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