

Mixed Gas Law Worksheet

Answers available at end

Name: _____ (show your work) period: _____

Convert:

$100 \text{ mmHg} = \text{_____ kPa}$

$300 \text{ mmHg} = \text{_____ atm}$

$120 \text{ atm} = \text{_____ mmHg}$

$300 \text{ torr} = \text{_____ kPa}$

$25 \text{ kPa} = \text{_____ mmHg}$

$25 \text{ mmHg} = \text{_____ atm}$

$300^\circ\text{C} = \text{_____ K}$

$150\text{K} = \text{_____}^\circ\text{C}$

$5000^\circ\text{C} = \text{_____ K}$

$300\text{K} = \text{_____}^\circ\text{C}$

$-255^\circ\text{C} = \text{_____ K}$

$3500\text{K} = \text{_____}^\circ\text{C}$

$200^\circ\text{C} = \text{_____}^\circ\text{F}$

$95^\circ\text{F} = \text{_____}^\circ\text{C} = \text{_____ K}$

$45^\circ\text{F} = \text{_____}^\circ\text{C} = \text{_____ K}$

- 1) A 70. liter sample of gas initially at 280°C is allowed to cool at constant pressure, what will the new volume be at 0°C ?

$P_1 = \text{_____} \quad P_2 = \text{_____}$

$V_1 = \text{_____} \quad V_2 = \text{_____}$

$T_1 = \text{_____} \quad T_2 = \text{_____}$

- 2) A 4.0 liter sample of gas initially at 25°C is allowed to cool at constant pressure, what will the new volume be at -5°C ?

$P_1 = \text{_____} \quad P_2 = \text{_____}$

$V_1 = \text{_____} \quad V_2 = \text{_____}$

$T_1 = \text{_____} \quad T_2 = \text{_____}$

- 3) A buoyancy vest (used in scuba diving) is filled to a volume of 2.3 liters at a pressure of 1 atmosphere (at the surface where the temperature is a balmy 300K), what will the volume of air be in the vest when the diver dives to a depth of about 90 feet where the temperature is 17°C and the pressure is 4 atmospheres?

$P_1 = \text{_____} \quad P_2 = \text{_____}$

Mixed Gas Law Review

S Ashworth



Mixed Gas Law Review:

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