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Mixed Stoichiometry Problems

Answer the following questions on your own piece of paper. You need to show all work for full credit. You will need to write a balanced equation to begin, and remember to include significant figures in your final answer.

1. Nitrogen gas reacts with hydrogen gas to produce ammonia (NH_3) gas. What volume of hydrogen is necessary to react with 5.00 L of nitrogen to produce ammonia?
2. What volume of ammonia is produced in the reaction from question one?
3. When propane (C_3H_8) is burned in oxygen, carbon dioxide and water are produced. If 20.0 L of oxygen are consumed in the above reaction, how many liters of carbon dioxide are produced?
4. In a chemical reaction, water is broken down into hydrogen and oxygen gases. If 30.0 mL of hydrogen are produced in this reaction, how many milliliters of oxygen are produced?
5. When carbon monoxide is burned in oxygen, carbon dioxide is produced. How many liters of carbon dioxide are produced if 75.0 L of carbon monoxide are burned in oxygen? How many liters of oxygen are necessary?
6. If ammonia (NH_3) was created by combining nitrogen and hydrogen, what volume of NH_3 at STP is produced if 25.0 g of N_2 is reacted with excess hydrogen?
7. Potassium chloride and oxygen are the products of the decomposition of potassium chlorate. If 5.0 g of potassium chlorate are decomposed, what volume of oxygen is produced at STP?
8. How many grams of potassium chloride are produced in problem 7?

Mixed Stoichiometry Word Problems

SJ Ball



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Stanford Bulletin, 2003

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