

## Solubility Rules for Common Ionic Compounds in Water

### Some Soluble Compounds

Group IA cations ( $\text{Li}^+$ ,  $\text{Na}^+$ ,  $\text{K}^+$ ,  $\text{Rb}^+$ ,  $\text{Cs}^+$ ),  $\text{NH}_4^+$

$\text{NO}_3^-$ ,  $\text{ClO}_4^-$ ,  $\text{CH}_3\text{COO}^-$ ,  $\text{ClO}_3^-$

$\text{Cl}^-$ ,  $\text{Br}^-$ ,  $\text{I}^-$

$\text{SO}_4^{2-}$

### Exceptions

No Exceptions

No Exceptions

Halides of  $\text{Ag}^+$ ,  $\text{Hg}_2^{2+}$ ,  $\text{Pb}^{2+}$

Sulfates of  $\text{Sr}^{2+}$ ,  $\text{Ba}^{2+}$ ,  $\text{Hg}_2^{2+}$ ,  $\text{Pb}^{2+}$

### Some Insoluble Compounds

$\text{S}^{2-}$

Sulfides of Group 1A cations,  $\text{NH}_4^+$ ,  $\text{Ca}^{2+}$ ,  $\text{Sr}^{2+}$ ,  $\text{Ba}^{2+}$

$\text{CO}_3^{2-}$

Carbonates of Group 1A cations and  $\text{NH}_4^+$

$\text{OH}^-$

Hydroxides of Group 1A cations,  $\text{NH}_4^+$ ,  $\text{Ca}^{2+}$ ,  $\text{Sr}^{2+}$ ,  $\text{Ba}^{2+}$

$\text{PO}_4^{3-}$

Phosphates of Group 1A cations and  $\text{NH}_4^+$

# Model Rules Of Solubility In Aqueous Solutions

**KJ Lindholm-Leary**



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