

CHAPTER 4

4.1 Chemical Reactions

Chemical Change

- the transformation of one or more substances into different substances with different properties

Chemical Reactions

- include a change in energy → energy is absorbed or released (heat, light, sound)
- occur at different rates: some are fast others are slow (a catalyst can be used to speed up reactions)

Clues to a chemical change

- Change in colour
- Energy is released or absorbed
- Gas is produced
- A precipitate (solid) forms from two liquids
- Difficult to reverse

NOTE: a change in state or dissolving a substance is NOT a chemical change

Chemical Equations

- Represent chemical reactions by using words or symbols and formulas
- Starting materials are called "reactants"
- New substances are called "products"
- An arrow (→) is read as "produces"
- States of matter appear in brackets after a chemical formula
- (s)=solid (l)=liquid (g)=gas (aq) dissolved in water

Word Equations

Ex. 1

- $\text{Mg (s)} + \text{HCl (aq)} \rightarrow \text{MgCl}_2 \text{ (aq)} + \text{H}_2 \text{ (g)}$
- Magnesium metal "reacts with" hydrochloric acid "to produce" Magnesium chloride and hydrogen gas

Ex. 2

- $\text{FeSO}_4 \text{ (aq)} + \text{NaCl} \rightarrow \text{FeCl}_2 \text{ (s)} + \text{Na}_2\text{SO}_4 \text{ (aq)}$
- Iron (ii) sulfate reacts with sodium chloride to produce iron (ii) chloride and sodium sulfate

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