

Chemistry – Unit 5 Worksheet 1

Atomic Structure of the Elements

Exercise 1: Mass From Counts

By a better understanding of atomic structure, we are able to understand the relationship between mass and structure.

1. Equal numbers of atoms contain the same number of nucleons as the same mass of an element. **Example: Argon and Oxygen**
2. Argon has a relative atomic mass of 39.948 and oxygen has a relative atomic mass of 15.999.

Example: $20 \times 39.948 = 798.96$ g. Argon has a relative atomic mass of 39.948 and oxygen has a relative atomic mass of 15.999. If we have 798.96 g of oxygen, we have 20 moles of oxygen. **Mass of 20 moles**

Argon
20 g Argon
1 mole

Oxygen
15.999 g
1 mole



Example:

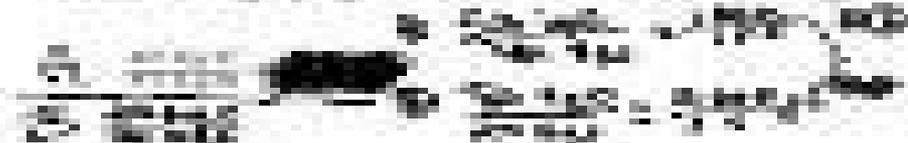
Oxygen = $\frac{15.999 \text{ g}}{1 \text{ mole}}$
 $\frac{798.96 \text{ g}}{15.999 \text{ g}} = 49.948 \text{ moles}$
 $49.948 \text{ moles} \times 39.948 \text{ g} = 1995.96 \text{ g}$

Chemists use the relative atomic mass of an element to determine the mass of all the elements using this technique. Measurements of the density of the gases of all elements have shown that the relative atomic mass of an element is the same as the mass of one mole of the gas. This is the same as the mass of one mole of the element in its standard state.

Exercise 2: Mass From Counts

Many elements are made up of different types of atoms called isotopes. In the periodic table you can find the relative atomic mass of an element. This is the average mass of the atoms of the element. The relative atomic mass of an element is the same as the mass of one mole of the element.

For example, the relative atomic mass of carbon is 12.011. This means that 12.011 g of carbon contains one mole of carbon atoms. The relative atomic mass of oxygen is 15.999. This means that 15.999 g of oxygen contains one mole of oxygen atoms.



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B Lingard

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