

## Unit 6 – Worksheet 3

### Ionic Compounds

#### Properties

#### Basic structural unit

1. Give the name of the following simple binary ionic compounds.

- $\text{Na}_2\text{O}$
- $\text{K}_2\text{S}$
- $\text{MgCl}_2$
- $\text{CaBr}_2$
- $\text{BaI}_2$
- $\text{Al}_2\text{S}_3$
- $\text{CaBr}$
- $\text{AgF}$

2. Give the name of the following simple binary ionic compounds.

- $\text{Na}_3\text{N}$
- $\text{K}_2\text{O}$
- $\text{AgBr}$
- $\text{MgI}_2$
- $\text{SrO}$

3. Write the formula for the following binary ionic compounds.

- lithium bromide
- sodium iodide
- silver sulfide
- cesium oxide
- beryllium iodide
- barium hydride
- aluminum fluoride
- potassium oxide

4. Write the formula for these ionic substances.

- silver oxide
- aluminum sulfide
- sodium nitride
- barium chloride
- strontium hydride

# Modeling Chemistry U6 Test

**W. T. Carnall**



## **Modeling Chemistry U6 Test:**

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## **Table of Contents Modeling Chemistry U6 Test**

1. Understanding the eBook Modeling Chemistry U6 Test
  - The Rise of Digital Reading Modeling Chemistry U6 Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U6 Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry U6 Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U6 Test
  - Personalized Recommendations
  - Modeling Chemistry U6 Test User Reviews and Ratings
  - Modeling Chemistry U6 Test and Bestseller Lists
5. Accessing Modeling Chemistry U6 Test Free and Paid eBooks
  - Modeling Chemistry U6 Test Public Domain eBooks
  - Modeling Chemistry U6 Test eBook Subscription Services
  - Modeling Chemistry U6 Test Budget-Friendly Options
6. Navigating Modeling Chemistry U6 Test eBook Formats

- ePub, PDF, MOBI, and More
  - Modeling Chemistry U6 Test Compatibility with Devices
  - Modeling Chemistry U6 Test Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Modeling Chemistry U6 Test
    - Highlighting and Note-Taking Modeling Chemistry U6 Test
    - Interactive Elements Modeling Chemistry U6 Test
  8. Staying Engaged with Modeling Chemistry U6 Test
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Modeling Chemistry U6 Test
  9. Balancing eBooks and Physical Books Modeling Chemistry U6 Test
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Modeling Chemistry U6 Test
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Modeling Chemistry U6 Test
    - Setting Reading Goals Modeling Chemistry U6 Test
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Modeling Chemistry U6 Test
    - Fact-Checking eBook Content of Modeling Chemistry U6 Test
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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