

## Unit 6 – Worksheet 3

### Ionic Compounds

#### Properties

#### Basic structural unit

1. Give the name of the following simple binary ionic compounds.

- $\text{Na}_2\text{O}$
- $\text{K}_2\text{S}$
- $\text{MgCl}_2$
- $\text{CaBr}_2$
- $\text{BaI}_2$
- $\text{Al}_2\text{S}_3$
- $\text{CaBr}$
- $\text{AgF}$

2. Give the name of the following simple binary ionic compounds.

- $\text{Na}_3\text{N}$
- $\text{K}_2\text{O}$
- $\text{AgBr}$
- $\text{MgI}_2$
- $\text{SrO}$

3. Write the formula for the following binary ionic compounds.

- lithium bromide
- sodium iodide
- silver sulfide
- cesium oxide
- beryllium iodide
- barium hydride
- aluminum fluoride
- potassium oxide

4. Write the formula for these ionic substances.

- silver oxide
- aluminum sulfide
- sodium nitride
- barium chloride
- strontium hydride

# Modeling Chemistry U6 Ws 2 V2 0

**JS Bruner**



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