

# Chemistry – Unit 7 Chemical Reactions

## Lesson 1: Atoms

### Background

Chemistry before 1900 generally focused on the study of elements. The scientific approach to study of the atom is recent.

### Historical View of Atomic Model of Matter

The atomic theory of matter postulates that all matter consists of very small particles which are constantly moving in a continuous motion. The discovery of the particulate nature demonstrated the physical energy that there are many similarities to other particles.

#### 1. Conservation of Mass

A principle stating that matter neither is created nor destroyed.

also known as the law of

#### 2. Chemical Formula

A set of chemical symbols representing atoms present in a compound and

their relative proportions, which define the structure of the compound.

#### 3. Molecular Mass

Relative to the atomic mass.

Number of atoms of an element present in any molecule or formula unit.

also

Chemical formula which shows the number of atoms of each element in a molecule.

A molecule is formed that characterizes a property.

also

#### 4. Pressure

1. The force exerted per unit area by the particles of the substance contained in a definite volume. This amount of force exerted by the particles contained in unit area is called pressure. It is defined as force applied per unit area perpendicular to the surface.

2. The force exerted by particles in a definite volume of a substance contained in a definite volume is called the pressure.

3. Chemists use number of atoms molecules present in a unit to study the properties of a substance. It is called the atomic mass, molecular mass, etc.

4. Chemists use number molecules present in a unit to study the properties of a substance. It is called the atomic mass, molecular mass, etc.

# Modeling Chemistry Unit 7 Rearranging Atoms Key

**Raffaella Di Napoli**



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