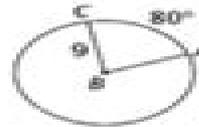


**Geometry**  
**Practice Test - Circles**

Name \_\_\_\_\_

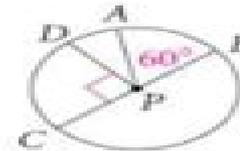
For Questions 1-2, leave your answers in terms of  $\pi$ .

- Find the length of  $\overline{CA}$  \_\_\_\_\_
- Find the area of sector CBA. \_\_\_\_\_



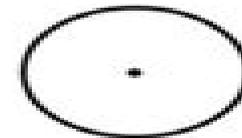
Find each measure for circle P.

- $m$  arc  $AB$  \_\_\_\_\_
- $m$  arc  $ABC$  \_\_\_\_\_
- $m \angle DPA$  \_\_\_\_\_
- $m$  arc  $DBC$  \_\_\_\_\_



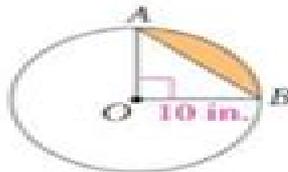
- Draw the following in the diagram below and label accordingly.

- Draw secant AB
- Draw tangent CD
- Draw chord EF



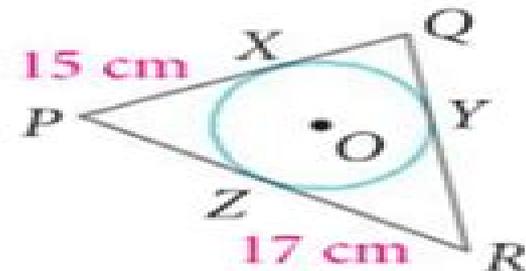
- Find the area of the shaded segment to the nearest tenth.

Area = \_\_\_\_\_

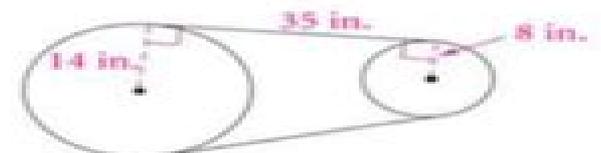


- Circle O is inscribed in  $\triangle PQR$ .  $\triangle PQR$  has a perimeter of 88 cm. Find QY.

QY = \_\_\_\_\_



- A belt fits tightly around two circular pulleys, as shown to the right. Find the distance between the centers of the pulleys.



# Geometry Circles Practice Test

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web cinderella here kitty kitty kitty kitty the cat yawns come kitty c mon the cat turns around not paying cinderella any mind lucifer come here lucifer begrudgingly walks out the door and yawns as cinderella closes the door on him i m sorry if your highness objects to an early breakfast

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web script print audition central disney s cinderella kids script stepsisters side 1 the narrators exit as cinderella enters balancing a huge breakfast tray on her head wincing from the bloodcurdling screams of her stepmother and stepsisters anastasia where s my muffin drizella where s my tea

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web girls don t forget to give cinderella your laundry it s dirty sock day the stepsisters dump a pile of dirty socks on cinderella s head and exit the mice clench their paws in anger

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web cinderella had an evil stepmother and two evil stepsisters named anastasia and drizella narrator 4 cold cruel and jealous they made cinderella a servant in her own house

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web voila finally the cinderella story script is here for all you quotes spouting fans of the movie starring hilary duff this script is a transcript that was painstakingly transcribed using the screenplay and or viewings of cinderella story i know i know i still

need to get the cast names in there and i ll be eternally tweaking it so if you have any corrections feel

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the ion meta dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be

conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab questions write the balanced equation

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate  $\text{CuSO}_4$  and 2 0153 grams of iron powder  $\text{Fe}$  the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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iron with copper ii sulfate i obtained the results to this experiment through adding heated  $\text{CuSO}_4$  solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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