

Chapter 7 Answers

Practice 7-1

1. 8 2. 8 3. 60 4. 240 5. 2,635 6. 1.92
 7. 0.8125 8. 9 9. 1500 10. 8.75 11. 3.44
 12. 7 13. 110 14. 30 15. 16 16. 119
 17. 12 18. 48 19. 40 20. 56

Practice 7-2

1. $\sqrt{51}$ 2. $4\sqrt{7}$ 3. $2\sqrt{65}$ 4. $\sqrt{7}$
 5. $2\sqrt{21}$ 6. $18\sqrt{2}$ 7. 210 in.^2 8. $24\sqrt{10} \text{ cm}^2$
 9. $35\sqrt{114} \text{ ft}^2$ 10. 46 in. 11. 78 ft 12. 279 cm
 13. 19 m 14. acute 15. right 16. obtuse
 17. right 18. obtuse 19. acute

Practice 7-3

1. $x = 2; y = \sqrt{3}$ 2. $a = 4.5; b = 4.5\sqrt{3}$
 3. $c = \frac{20\sqrt{3}}{3}; d = \frac{20\sqrt{3}}{3}$ 4. $8\sqrt{2}$ 5. $14\sqrt{2}$
 6. $\frac{20\sqrt{3}}{3}$ 7. 2 8. $x = 15; y = 15\sqrt{3}$ 9. $3\sqrt{2}$
 10. 42 cm 11. 5.9 in. 12. 10.4 ft, 12 ft 13. 170 in.²
 14. $w = \frac{20\sqrt{3}}{3}; x = 5; y = 5\sqrt{2}; z = \frac{2\sqrt{3}}{3}$ 15. $a = 4;$
 $b = 3$ 16. $p = 4\sqrt{3}; q = 4\sqrt{3}; r = 8; s = 4\sqrt{6}$

Practice 7-4

1. 48 cm^2 2. 784 in.^2 3. 11.4 ft^2 4. 90 m^2
 5. 2400 in.^2 6. 374 ft^2 7. 160 cm^2 8. 176.25 in.^2
 9. 54 ft^2 10. 42.5 square units 11. $96\sqrt{3}$ square units
 12. $36\sqrt{3}$ square units 13. 45 cm^2 14. 226.2 in.^2
 15. $49,500 \text{ m}^2$

Practice 7-5

1. $x = 7; y = 3.5\sqrt{3}$ 2. $a = 60; c = \frac{8\sqrt{3}}{3}; d = \frac{8\sqrt{3}}{3}$
 3. $p = 4\sqrt{3}; q = 8; s = 30$ 4. $m\angle 1 = 45; m\angle 2 = 45;$
 $m\angle 3 = 90; m\angle 4 = 90$ 5. $m\angle 5 = 60; m\angle 6 = 30;$
 $m\angle 7 = 120$ 6. $m\angle 8 = 60; m\angle 9 = 30; m\angle 10 = 60$
 7. $12\sqrt{3}$ 8. $36.75\sqrt{3}$ 9. $75\sqrt{3}$ 10. 120 in.^2
 11. 137 in.^2 12. 97 in.^2

Practice 7-6

1. 32π 2. 16π 3. 7.8π 4. Samples: $\overline{DFA}, \overline{ABF}$
 5. Samples: $\overline{FE}, \overline{FD}$ 6. Samples: $\overline{FEB}, \overline{EDA}$
 7. Samples: $\overline{FE}, \overline{ED}$ 8. $\angle ACB$ 9. $\angle FCD$
 10. 32 11. 43 12. 54 13. 39 14. 71
 15. 90 16. 50 17. 220 18. 140 19. 220
 20. 130 21. $6\pi \text{ in.}$ 22. $16\pi \text{ cm}$ 23. $\frac{2}{3}\pi \text{ m}$

Practice 7-7

1. 49π 2. 21.2 3. $\frac{20}{3}\pi$ 4. $\frac{20}{3}\pi - 21.2$ 5. $\frac{1}{3}\pi$
 or $\frac{\pi}{3}$ 6. $\frac{1}{3}$ 7. $\frac{1}{10}\pi$ or $\frac{\pi}{10}$ 8. $\frac{\pi}{10} - \frac{1}{3}$ 9. 4π
 10. $\frac{10}{3}\pi$ 11. $\frac{2}{3}\pi$ 12. $\frac{2}{3}\pi$ 13. $\frac{10}{3}\pi$ 14. 6π
 15. $\frac{20}{3}\pi$ 16. $\frac{20}{3}\pi$ 17. $\frac{\pi}{3}$ 18. 3 19. 2
 20. 10

Practice 7-8

1. 8.7% 2. 10.9% 3. 15.3% 4. $\frac{3}{10}$ 5. $\frac{7}{10}$
 6. $\frac{3}{5}$ 7. $\frac{1}{10}$ 8. 1 9. $\frac{1}{2}$ 10. 22.2% 11. 4.0%
 12. 37.5% 13. 20% 14. 50% 15. 40%
 16. $33\frac{1}{3}\%$

Reteaching 7-1

1. Samples:



2. Samples:



3. 20 cm^2 4. 55 in.^2 5. 6 m 6. 8 ft 7. 4 in.

8. Samples:



9. 5 in. 10. 23 cm 11. 9 ft

Reteaching 7-2

1.–2. Sample:



3. Sample: 9, 16, and 25 4. They are equal. 5. The sum of the squares of the two legs equals the square of the hypotenuse. 6. $4\sqrt{5}$ 7. $7\sqrt{2}$ 8. 8 9. 6

Geometry Lesson Master Answer Key 10 3a

SJ Ball



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web $m \times c \times \Delta t$ equation 1 where m is the total mass of the solution solute plus solvent c is the specific heat of the solution and Δt is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of *guided inquiry designing the ideal hand warmer - Dec 27 2022*

web the heat of reaction ΔH_{soln} is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$ $\Delta H_{\text{soln}} = 44.2 \text{ kJ mol rxn}$ write equations to similarly represent the dissolving process for

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web the goal of my experiment and research is to evaluate the various number of compounds to be used as hand warmers and locate the ones that can raise the temperature to 20 degrees celsius in 100 g water and costs as little as possible

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution ΔH_{soln} or the amount of energy released or absorbed in kJ/mol of the substance being dissolved

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