

**Mixed Mole Problems Worksheet**

Solve the following problems. Be sure to show all your work!

1. Find the number of moles of silver ions (
- $Ag^+$
- ) or
- $Ag^+$
- ions in a solution of silver nitrate.

$$1) \quad 1.5 \text{ L of } 0.10 \text{ M } AgNO_3 \text{ containing } Ag^+ \quad 1.5 \text{ L} \times 0.10 \text{ M } AgNO_3 = \boxed{0.15 \text{ mol } Ag^+}$$

2. How many moles of
- $CaCl_2$
- are there in 2.00 kg of
- $CaCl_2$
- ?

$$2) \quad 2.00 \text{ kg } CaCl_2 \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mol } CaCl_2}{110.98 \text{ g}} = \boxed{18.0 \text{ mol } CaCl_2}$$

3. What is the mass of 0.025 moles of
- $CaCl_2$
- ? (molar mass of
- $CaCl_2 = 110.98 \text{ g/mol}$
- )

$$3) \quad 0.025 \text{ mol } CaCl_2 \times 110.98 \text{ g/mol} = \boxed{2.77 \text{ g } CaCl_2}$$

4. Find the number of moles of ions in 2.00 g of
- $CaCl_2$
- (molar mass of
- $CaCl_2 = 110.98 \text{ g/mol}$
- )

$$4) \quad 2.00 \text{ g } CaCl_2 \times \frac{1 \text{ mol } CaCl_2}{110.98 \text{ g}} \times \frac{3 \text{ mol ions}}{1 \text{ mol } CaCl_2} = \boxed{0.053 \text{ mol ions}}$$

5. If you start off with 0.025 moles of
- $CaCl_2$
- , what is the molar mass of this salt? Please show all your work.

$$\frac{0.27 \text{ g}}{0.025 \text{ mol}} = \frac{M}{0.025 \text{ mol}} \quad \begin{array}{l} \text{molar mass} \\ \text{of } CaCl_2 \end{array} \quad \frac{0.27 \text{ g}}{0.025 \text{ mol}} = \frac{M}{0.025 \text{ mol}} \quad \text{molar mass of } CaCl_2$$

$$\boxed{M = 10.8 \text{ g/mol}}$$

6. What volume of 0.10 M
- $CaCl_2$
- solution contains 0.025 moles of
- $CaCl_2$
- ?

$$6) \quad 0.025 \text{ mol } CaCl_2 \times \frac{1 \text{ L solution}}{0.10 \text{ mol } CaCl_2} = \frac{0.025 \text{ mol}}{0.10 \text{ mol/L}} = \boxed{0.25 \text{ L of } CaCl_2}$$

7. What is the mass of 0.025 moles of
- $CaCl_2$
- ? (molar mass of
- $CaCl_2 = 110.98 \text{ g/mol}$
- )

$$7) \quad 0.025 \text{ mol } CaCl_2 \times \frac{110.98 \text{ g}}{1 \text{ mol } CaCl_2} = \frac{2.77 \text{ g } CaCl_2}{1 \text{ mol } CaCl_2} = \boxed{2.77 \text{ g } CaCl_2}$$

8. Calculate the molar mass of a salt if 0.025 moles of
- $CaCl_2$
- is 0.27 g. What is the molar mass of
- $CaCl_2$
- ?

$$\frac{0.27 \text{ g}}{0.025 \text{ mol}} = \frac{M}{0.025 \text{ mol}} \quad \begin{array}{l} \text{molar mass of } CaCl_2 \\ \text{salt} \end{array} \quad \frac{0.27 \text{ g}}{0.025 \text{ mol}} = \frac{M}{0.025 \text{ mol}} \quad \text{molar mass of } CaCl_2$$

# Mixed Mole Problems Answers Show Work

**JA Banks**



**Mixed Mole Problems Answers Show Work:**

## Whispering the Techniques of Language: An Emotional Quest through **Mixed Mole Problems Answers Show Work**

In a digitally-driven earth where monitors reign supreme and quick communication drowns out the subtleties of language, the profound strategies and psychological subtleties hidden within phrases frequently go unheard. Yet, nestled within the pages of **Mixed Mole Problems Answers Show Work** a captivating literary value pulsating with raw emotions, lies a fantastic quest waiting to be undertaken. Written by a talented wordsmith, this enchanting opus invites viewers on an introspective journey, delicately unraveling the veiled truths and profound influence resonating within the material of each word. Within the mental depths with this touching evaluation, we will embark upon a sincere exploration of the book's key styles, dissect their interesting writing model, and succumb to the strong resonance it evokes strong within the recesses of readers' hearts.

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