

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

- Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

Equation:	<u> 2 </u> H ₂ S _(g) + <u> 3 </u> O _{2(g)} → <u> 2 </u> SO _{2(g)} + <u> 2 </u> H ₂ O _(g)
Before	<u> 2 </u> <u> ex </u> <u> 0 </u> <u> 0 </u> 3.0 mol O ₂
Change	<u> -2 </u> <u> -3 </u> <u> 2 </u> <u> 2 </u>
After	<u> 0 </u> <u> ex </u> <u> 2 </u> <u> 2 </u>

- Propane, C₃H₈, burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?

Equation: $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$

Before	0.5	ex	0	0	1.50mol HCL
Change	-0.5	-1.5	+0.5	+0.5	
After	0	ex	0.5	0.75	

- Ammonia, NH₃, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?

Equation: $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$

Before	ex	1.25	0	2.50 mole
Change	-3.75	-1.25	+25	
After	ex	0	2.5	

Modeling Chemistry Unit 8

JS Bruner



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