

EQUILIBRIUM CONSTANT - WORKSHEET 2_ANSWERS

1. Write the equilibrium constant expression for each of the following reactions:



$$K = \frac{[\text{NO}_2]^2}{[\text{N}_2][\text{O}_2]}$$



$$K = \frac{[\text{C}_2\text{H}_6]}{[\text{H}_2]^3}$$



$$K = \frac{1}{[\text{I}_2(\text{aq})]}$$

2. For the following reaction:



At equilibrium, $[\text{Fe}^{2+}] = 0.50\text{ M}$, $[\text{Ag}^+] = 1.0\text{ M}$, $[\text{Fe}^{3+}] = 1.50\text{ M}$

a) Calculate the equilibrium constant for the reaction.

$$K = 3.0$$

b) If at equilibrium, $[\text{Fe}^{2+}] = 0.20\text{ M}$ and $[\text{Ag}^+] = 0.30\text{ M}$, what must be the concentration of Fe^{3+} ?

$$[\text{Fe}^{3+}] = 0.18\text{ mol/L}$$



Calculate the equilibrium concentration of oxygen gas if the NO and NO_2 concentrations are equal at equilibrium.

$$[\text{O}_2] = 2.0 \times 10^{-6}\text{ mol/L}$$



Which of the above reactions would have the greatest amount of products present at equilibrium compared to reactants?

Reaction b. Since K is larger the amount of products must be larger, because K is products/reactants.

5. For the reaction:



at an equilibrium temperature of 1000°C , a 1.00 L flask contains 0.120 moles of ammonia, 1.03 moles of nitrogen and 1.62 moles of hydrogen. Calculate the equilibrium constant at this temperature.

$$K = 3.29 \times 10^{-2}$$

Genetic Equilibrium Worksheet Answer Key

David Kirk



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