

Holt Modern Chemistry Review

CHAPTER 5: THE PERIODIC LAW

The following pages contain the bulk (but not all) of the information for the chapter 5 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and **NOTHING** else, you should at least be able to **PASS** the test. ***** Test items will be recall, examples, and/or application of this content. *****

OUTCOMES

- Collaborate with peer(s) to understand chemistry content (**C C**)
- Communicate chemistry content to teacher and peer(s) (**E C**)
- 5.1: Describe the organization of the modern Periodic table; distinguish between periods and groups (**T & R**)
- 5.2: Use the Periodic Table to obtain information about the properties (physical and chemical), ionization energy, and atomic sizes of the elements (**F & PK**)
- 5.2: Identify common non-metals, metalloids, and noble gases; identify the representative groups and name them (groups 1, 2, 13, 14, 15, 16, 17, 18) (**T & R**)
- 5.3: Compare periodic trends of atomic radii, ionization energy, and electronegativity (**T & R**)

5.1: HISTORY OF THE PERIODIC TABLE

- **Vocabulary**
 - **periodic law** -- the law that states that the repeating chemical and physical properties of elements change periodically with the atomic numbers of the elements
 - **periodic table** -- an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column
 - **lanthanides** -- a member of the rare-earth series of elements, whose atomic numbers range from 58 (cerium) to 71 (lutetium)
 - **actinides** -- any of the series of heavy radioactive elements that extends from thorium (atomic number 90) through lawrencium (atomic number 103) on the periodic table
- **Chapter Highlights**
 - The periodic law states that the physical and chemical properties of the elements are periodic functions of their atomic numbers.
 - The periodic table is an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column/group/family.
 - The columns in the periodic table are referred to as groups.

5.2: ELECTRON CONFIGURATION AND THE PERIODIC TABLE

- **Vocabulary**
 - **alkali metal** -- one of the elements of Group 1 of the periodic table (lithium, sodium, potassium, rubidium, cesium, and francium)
 - **alkaline-earth metal** -- one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, barium, and radium)
 - **transition element** -- one of the metals that can use the inner shell before using the outer shell to bond
 - **main-group element** -- an element in the s-block or p-block of the periodic table
 - **halogen** -- one of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine); halogens combine with most metals to form salts
- **Chapter Highlights**
 - The rows in the periodic table are called periods.
 - Many chemical properties of the elements can be explained by the configurations of the elements' outermost (valence) electrons.
 - The noble gases exhibit unique chemical stability because their highest occupied levels have an octet (8) of electrons, $?s^2?p^6$ (with the exception of helium, whose stability arises from its highest occupied level being completely filled with two electrons, $1s^2$).

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Jin-Ying Zhang



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