

# The Mole Concept: Chemical Formula of a Hydrate – Lab Report Assistant

## Exercise 1: Water of Hydration

Data Table 1. Alum Data.

Object	Mass (g)
Aluminum Cup (Empty)	2.5 (g)
Aluminum Cup + 2.00 grams of Alum	4.5 (g)
Aluminum Cup + Alum After 1 <sup>st</sup> Heating	3.6 (g)
Aluminum Cup + Alum After 2 <sup>nd</sup> Heating	3.5 (g)
Mass of Released H <sub>2</sub> O	1.0 (g)
Molecular Mass of H <sub>2</sub> O	18 g/mol
Moles of Released H <sub>2</sub> O	0.0556 mol

### Questions:

- A) Calculate the moles of anhydrous (dry)  $KAl(SO_4)_3$ , that were present in the sample. Show all work including units.
- B) Calculate the ratio of moles of  $H_2O$  to moles of anhydrous  $KAl(SO_4)_3$ . Show all work including units. *Note: Report the ratio to the closest whole number.*
- C) Write the empirical formula for the hydrated  $KAl(SO_4)_3$ , based on your experimental results and answer to Question 2. Show all work including units. *Hint: if the ratio of moles of  $H_2O$  to moles of anhydrous  $KAl(SO_4)_3$  was 4, then the empirical formula would be:  $KAl(SO_4)_3 \cdot 4H_2O$ .*
- D) Describe any visual differences between the hydrated sample and the dried, anhydrous form.
- E) How would the following errors affect the empirical formula for the compound? That is, will these errors cause the calculated number of moles of water in the hydrate to be artificially high or low?
  - a. The student ran out of time and did not do the second heating. Explain how this error will affect the calculation for the number of moles of water in the hydrate? Will the final answer be artificially high or low? How do you know?
  - b. The student recorded the mass of the cup + sample incorrectly and started with 2.20 g of hydrated compound but used 2.00 g in the calculations. Explain how this error will affect the calculation for the number of moles of water in the hydrate? Will the final answer be artificially high or low? How do you know?

# Hydrates Lab Answers

**J Spring**



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