

Oxford Resources for IB

Chemistry – 2023 Edition

Answers

Reactivity 1.1 – Measuring enthalpy changes

Activity (page 390)

- The glass of water is an open system.
- The system includes the water and ice in the glass. The surroundings are all matter that exists outside the glass of iced water. The glass can be considered the boundary between the system and the surroundings.
- Heat, q , is a form of energy that is transferred from a warmer body to a cooler body, as a result of the temperature gradient. The water in the glass is warmer than the ice cubes ($\text{H}_2\text{O}(\text{s})$). Energy will flow from the water into the ice cubes, warming them, so eventually they will melt. As the surroundings are warmer than the iced water, energy will flow from the surroundings into the system.
- Condensation occurs on the outside of the glass and you observe liquid water. The change of state is from a gas (water vapour) to a liquid (water). Energy is transferred into the surroundings when water vapour condenses into liquid water.

Practice question (page 393)

1. C

Practice questions (page 394)

2. a. $q = mc\Delta T$
 $= 1 \text{ kg} \times 4.18 \text{ kJ kg}^{-1} \text{ K}^{-1} \times 1 \text{ K}$
 $= 4.18 \text{ kJ}$
- b. $q = mc\Delta T$
 $= 1000 \text{ kg} \times 0.385 \text{ kJ kg}^{-1} \text{ K}^{-1} \times 1 \text{ K}$
 $= 385 \text{ kJ}$

3. a

4. a

Practice questions (page 396)

5. $q = mc\Delta T$
 $= 30 \text{ g} \times 4.18 \text{ J g}^{-1} \text{ K}^{-1} \times 30 \text{ K}$
 $= 3762 \text{ J}$
6. First, convert kJ to J ($\times 1000$)
Next, convert g to kg ($\div 1000$)
 $q = mc\Delta T$
 $\Delta T = \frac{q}{mc}$
 $= \frac{675 \text{ J}}{0.125 \text{ kg} \times 385 \text{ J kg}^{-1} \text{ K}^{-1}}$
 $= 14.0 \text{ K (3 sf)}$



Ib Chemistry 2012 Paper Answers

J Elliott



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