

1. A certain gas obeys the equation of state $U(S, V, N) = aS^{\gamma}/V^{\delta}N^2$, where a is a dimensioned constant. Here U represents the internal energy of the gas, S the entropy, V the volume and N the fixed number of particles of the system.

- (a) Let such a gas be filled in a box of volume V and the internal energy of the system be U . A partition is placed to divide the box into two equal parts, each having volume $V/2$. For each part, the internal energy is now αU and the dimensioned constant be βa . Obtain α and β . [3]

$\alpha =$	$\beta =$
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- (b) The temperature T can be expressed in terms of the derivative of internal energy as [2]

$$T = \left(\frac{dU}{dS} \right)_{V, N}$$

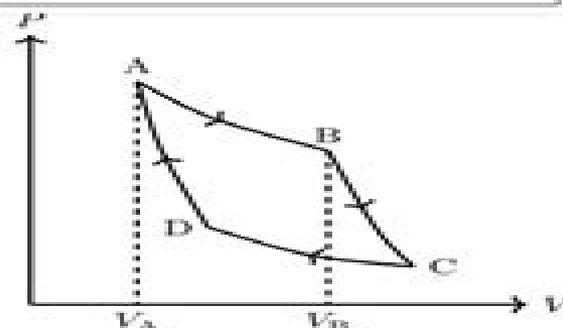
where the subscripts indicate that the differentiation has been carried out keeping V and N constant. In a similar way, express pressure P in terms of a derivative of the internal energy.

$P =$

- (c) Find the equation of state of the given system relating P , T , and V . [1]

$P =$

- (d) One mole of this gas executes a Carnot cycle ABCDA between reservoirs at temperatures T_1 and T_2 ($T_1 > T_2$). Obtain the heat change in the process AB (Q_{AB}) and work done by the system in the processes AB and BC (W_{AB} , W_{BC}) of the cycle. Express your answers only in terms of temperatures T_1 , T_2 , volumes V_A , V_B , and the other constants. [7]



$Q_{AB} =$

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Table of Contents International Physics Olympiad Question Paper With Solution

1. Understanding the eBook International Physics Olympiad Question Paper With Solution
 - The Rise of Digital Reading International Physics Olympiad Question Paper With Solution
 - Advantages of eBooks Over Traditional Books
2. Identifying International Physics Olympiad Question Paper With Solution
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an International Physics Olympiad Question Paper With Solution
 - User-Friendly Interface
4. Exploring eBook Recommendations from International Physics Olympiad Question Paper With Solution
 - Personalized Recommendations
 - International Physics Olympiad Question Paper With Solution User Reviews and Ratings
 - International Physics Olympiad Question Paper With Solution and Bestseller Lists

5. Accessing International Physics Olympiad Question Paper With Solution Free and Paid eBooks
 - International Physics Olympiad Question Paper With Solution Public Domain eBooks
 - International Physics Olympiad Question Paper With Solution eBook Subscription Services
 - International Physics Olympiad Question Paper With Solution Budget-Friendly Options
6. Navigating International Physics Olympiad Question Paper With Solution eBook Formats
 - ePub, PDF, MOBI, and More
 - International Physics Olympiad Question Paper With Solution Compatibility with Devices
 - International Physics Olympiad Question Paper With Solution Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of International Physics Olympiad Question Paper With Solution
 - Highlighting and Note-Taking International Physics Olympiad Question Paper With Solution
 - Interactive Elements International Physics Olympiad Question Paper With Solution
8. Staying Engaged with International Physics Olympiad Question Paper With Solution
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers International Physics Olympiad Question Paper With Solution
9. Balancing eBooks and Physical Books International Physics Olympiad Question Paper With Solution
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection International Physics Olympiad Question Paper With Solution
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine International Physics Olympiad Question Paper With Solution
 - Setting Reading Goals International Physics Olympiad Question Paper With Solution
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of International Physics Olympiad Question Paper With Solution
 - Fact-Checking eBook Content of International Physics Olympiad Question Paper With Solution
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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