

### Lab 5: Equilibrium

In the equilibrium lab, the equilibrium constant for the formation of iron(II) thiocyanate was spectroscopically determined. In a similar experiment, the formation constant of  $\text{Co}(\text{SCN})_4^{2-}$  was determined by mixing solutions of  $\text{Co}(\text{NO}_3)_2$  and  $\text{KSCN}$  and measuring the absorbance of the blue  $\text{Co}(\text{SCN})_4^{2-}$  formed in the reaction.



(a) What is the equilibrium expression for the reaction under study?

$K = \frac{[\text{Co}(\text{SCN})_4^{2-}]}{[\text{Co}^{2+}] \cdot (4 \times [\text{SCN}^{-}])}$

$K = \frac{[\text{Co}(\text{SCN})_4^{2-}]}{[\text{Co}^{2+}] \cdot [\text{SCN}^{-}]^4}$

$K = \frac{[\text{Co}^{2+}] \cdot [\text{SCN}^{-}]^4}{[\text{Co}(\text{SCN})_4^{2-}]}$

$K = \frac{[\text{Co}(\text{SCN})_4^{2-}]}{[\text{Co}^{2+}] + (4 \times [\text{SCN}^{-}])}$

(b) Given the data in the table, please complete the chart below. The molar absorptivity of  $\text{Co}(\text{SCN})_4^{2-}$  is  $1800 \text{ M}^{-1} \text{ cm}^{-1}$  at 610 nm. Assume the Beer's Law plot has a slope of 1800, and a y-intercept of zero.

Experimental Results

Volume of $1.00 \times 10^{-2} \text{ M Co}^{2+}$ (mL)	Volume of $8.00 \times 10^{-2} \text{ M SCN}^{-}$ (mL)	Initial $[\text{Co}^{2+}]$ (M)	Initial $[\text{SCN}^{-}]$ (M)	Absorbance at 610 nm	Calculated Equilibrium $[\text{Co}(\text{SCN})_4^{2-}]$ (M)
4.00	6.00	<input type="text" value="0.00025"/>	<input type="text" value="0.0012"/>	0.115	<input type="text" value="0.00025"/>

(c) Complete the reaction table for these results. (All entries should be in molarity.)

	$\text{Co}^{2+}(\text{aq})$	$+ 4 \text{SCN}^{-}(\text{aq})$	$\rightleftharpoons \text{Co}(\text{SCN})_4^{2-}(\text{aq})$
[initial]	<input type="text" value="0.00025"/>	<input type="text" value="0.0012"/>	0
[Δ]	<input type="text" value="0.00025"/>	<input type="text" value="0.0012"/>	<input type="text" value="0.00025"/>
[final]	<input type="text" value="0.00025"/>	<input type="text" value="0.0012"/>	<input type="text" value="0.00025"/>

(d) Calculate the formation constant for the reaction. (Enter your answer to three significant figures.)

$K_f =$

# Introducing Equilibrium Lab Answers

**Michael Brown**



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## **Table of Contents Introducing Equilibrium Lab Answers**

1. Understanding the eBook **Introducing Equilibrium Lab Answers**
  - The Rise of Digital Reading **Introducing Equilibrium Lab Answers**
  - Advantages of eBooks Over Traditional Books
2. Identifying **Introducing Equilibrium Lab Answers**
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an **Introducing Equilibrium Lab Answers**
  - User-Friendly Interface
4. Exploring eBook Recommendations from **Introducing Equilibrium Lab Answers**
  - Personalized Recommendations
  - **Introducing Equilibrium Lab Answers** User Reviews and Ratings

- Introducing Equilibrium Lab Answers and Bestseller Lists
- 5. Accessing Introducing Equilibrium Lab Answers Free and Paid eBooks
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  - Introducing Equilibrium Lab Answers eBook Subscription Services
  - Introducing Equilibrium Lab Answers Budget-Friendly Options
- 6. Navigating Introducing Equilibrium Lab Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Introducing Equilibrium Lab Answers Compatibility with Devices
  - Introducing Equilibrium Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Introducing Equilibrium Lab Answers
  - Highlighting and Note-Taking Introducing Equilibrium Lab Answers
  - Interactive Elements Introducing Equilibrium Lab Answers
- 8. Staying Engaged with Introducing Equilibrium Lab Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Introducing Equilibrium Lab Answers
- 9. Balancing eBooks and Physical Books Introducing Equilibrium Lab Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Introducing Equilibrium Lab Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Introducing Equilibrium Lab Answers
  - Setting Reading Goals Introducing Equilibrium Lab Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Introducing Equilibrium Lab Answers
  - Fact-Checking eBook Content of Introducing Equilibrium Lab Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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