

## 5

## ELECTRONS IN ATOMS

**SECTION 5.1 MODELS OF THE ATOM (pages 127–132)**

*This section summarizes the development of atomic theory. It also explains the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom.*

**► The Development of Atomic Models (pages 127–128)**

1. Complete the table about atomic models and the scientists who developed them.

Scientist	Model of Atom
Dalton	The atom is a solid indivisible mass.
Thomson	The atom is a ball of positive charge with electrons stuck into the ball.
Rutherford	Most of an atom's mass is concentrated in the small, positively charged nucleus. The electrons surround the nucleus and the rest of the atom is empty space.
Bohr	Electrons are arranged in concentric circular paths around the nucleus.

2. Is the following sentence true or false? The electrons in an atom can exist between energy levels. false

**► The Bohr Model (pages 128–129)**

3. What is a small, discrete unit of energy called?  
It is called a quantum.
4. Circle the letter of the term that completes the sentence correctly: A quantum of energy is the amount of energy required to
- move an electron from its present energy level to the next lower one
  - maintain an electron in its present energy level
  - move an electron from its present energy level to the next higher one
5. In general, the higher the electron is on the energy ladder, the farther it is from the nucleus.

# Introduction To Atoms Guided Reading And Study Answers

**David Kirk**



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