

### Question 1

A 5.00 L sample of  $N_2(g)$  is held at a pressure of 2.00 atm and a temperature of 315 K. Assuming ideal-gas behavior, how many moles of  $N_2(g)$  are present in this sample?

- a. 2.69 mol
- b. 0.208 mol
- c. 0.387 mol
- d. 0.004 mol

### Question 2

At a temperature of  $500^\circ\text{C}$  and a pressure of 699 Torr, sulfur vapor has a density of  $3.71\text{ g/dm}^3$ . What molecular formula for sulfur is compatible with this set of conditions?

- a.  $S_2$
- b.  $S_8$
- c.  $S_4$
- d.  $(S_{25})_2$

### Question 3

Consider a gas for which the equation of state  $P(V_m - b) = RT$  (with  $b = 0.04\text{ dm}^3/\text{mol}$ ) can be used to describe its P-V-T behavior. Suppose a 2-mol sample of this gas is compressed isothermally and reversibly from an initial volume of  $10\text{ dm}^3$  to a final volume of  $1\text{ dm}^3$ . How much work ( $w$ ) is done on the gas sample in this compression process, and how much is the internal energy ( $U$ ) of the gas changed?

- a.  $w = 11.7\text{ kJ}$  and  $\Delta U = 11.7\text{ kJ}$
- b.  $w = 11.9\text{ kJ}$  and  $\Delta U = 0$
- c.  $w = 5.85\text{ kJ}$  and  $\Delta U = -5.85\text{ kJ}$
- d.  $w = 5.85\text{ kJ}$  and  $\Delta U = 0$

### Question 4

Consider a process in which a sample of water vapor expands reversibly and adiabatically from a state in which its pressure is 60 Torr and its volume is  $0.5\text{ dm}^3$  to a state in which its volume is  $3\text{ dm}^3$ . If it is assumed that the water vapor behaves as an ideal gas (with respect to its P-V-T properties) and that the ratio of its  $C_{m,m}$  and  $C_{m,v}$  heat-capacity quantities is given by  $\gamma = C_{m,m}/C_{m,v} = 1.3$ , what will be the final pressure of the vapor sample?

- a. 5.83 bar
- b. 230 Torr
- c. 5.84 Torr
- d. There is not enough information provided to answer this question.

### Question 5

Consider a process in which exactly 107 J of heat is added to 1 mol of  $CH_3OH(g)$  under constant-volume conditions. If the initial temperature of the methanol vapor sample is 298 K, what will be the temperature of the vapor after the 107 J of heat is added? You may assume ideal-gas behavior and use a value of  $43.89\text{ J/K}\cdot\text{mol}$  for  $C_{m,m}$  (the molar heat capacity of methanol vapor under constant-pressure conditions).

- a. 301.0 K
- b. 300.4 K
- c. 295.0 K
- d. None of the above answers is correct.

### Question 6

If a person breathes at a rate of 14 respirations per minute, and each respiration involves  $0.5\text{ dm}^3$  of air, what will be the total mass of air breathed in a day? (Take  $T = 300\text{ K}$  and  $P = 101.3\text{ kPa}$ , and assume that the air behaves as an ideal gas with a molar mass of  $0.029\text{ kg/mol}$ ).

- a. 11.9 kg
- b. 119 g
- c. 238 g
- d. 14.6 kg

### Question 7

If you have 1.55 mol of  $F_2(g)$  at 1.11 atm pressure and a temperature of  $0^\circ\text{C}$ , what is the volume of this gas sample, assuming ideal-gas behavior?

- a. 31.3 L
- b. 22.4 L
- c. 3170 L
- d. 0.0821 L

### Question 8

The density of ice,  $H_2O(s)$ , at 273 K, is  $0.915\text{ kg/dm}^3$ , and that of liquid water is  $0.99987\text{ kg/dm}^3$ . How much thermodynamic work must be done to melt 1 mol of ice at 1 bar pressure?

- a. 0.167 J
- b. 0.167 kJ
- c. 1.67 J
- d. 39.9 cal

### Question 9

The heat capacity of a substance can be measured under either constant-volume or constant-pressure conditions. Which of the following sets of relationships between  $C_v$  (heat capacity under constant-volume conditions) and  $C_p$  (heat capacity under constant-pressure conditions) are always true?

- a.  $C_p(\text{gas}) > C_v(\text{gas})$ ;  $C_p(\text{solid}) = C_v(\text{solid})$ ; and  $C_p(\text{gas}) >> C_p(\text{solid})$
- b.  $C_v(\text{gas}) = C_v(\text{solid})$  and  $C_p(\text{liquid}) = C_v(\text{liquid})$
- c.  $C_v(\text{gas}) > C_v(\text{liquid}) > C_v(\text{solid})$  and  $C_p > C_v$  in all phases
- d.  $C_p$  approaches a value of 0 as the sample temperature approaches 0 K but increases without bound as the sample temperature is increased.

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### **Table of Contents Multiple Choice Physical Chemistry Exam Questions**

1. Understanding the eBook Multiple Choice Physical Chemistry Exam Questions
  - The Rise of Digital Reading Multiple Choice Physical Chemistry Exam Questions
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Physical Chemistry Exam Questions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Physical Chemistry Exam Questions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Physical Chemistry Exam Questions
  - Personalized Recommendations

- Multiple Choice Physical Chemistry Exam Questions User Reviews and Ratings
- Multiple Choice Physical Chemistry Exam Questions and Bestseller Lists
- 5. Accessing Multiple Choice Physical Chemistry Exam Questions Free and Paid eBooks
  - Multiple Choice Physical Chemistry Exam Questions Public Domain eBooks
  - Multiple Choice Physical Chemistry Exam Questions eBook Subscription Services
  - Multiple Choice Physical Chemistry Exam Questions Budget-Friendly Options
- 6. Navigating Multiple Choice Physical Chemistry Exam Questions eBook Formats
  - ePub, PDF, MOBI, and More
  - Multiple Choice Physical Chemistry Exam Questions Compatibility with Devices
  - Multiple Choice Physical Chemistry Exam Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Multiple Choice Physical Chemistry Exam Questions
  - Highlighting and Note-Taking Multiple Choice Physical Chemistry Exam Questions
  - Interactive Elements Multiple Choice Physical Chemistry Exam Questions
- 8. Staying Engaged with Multiple Choice Physical Chemistry Exam Questions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Multiple Choice Physical Chemistry Exam Questions
- 9. Balancing eBooks and Physical Books Multiple Choice Physical Chemistry Exam Questions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Multiple Choice Physical Chemistry Exam Questions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Multiple Choice Physical Chemistry Exam Questions
  - Setting Reading Goals Multiple Choice Physical Chemistry Exam Questions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Multiple Choice Physical Chemistry Exam Questions
  - Fact-Checking eBook Content of Multiple Choice Physical Chemistry Exam Questions

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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