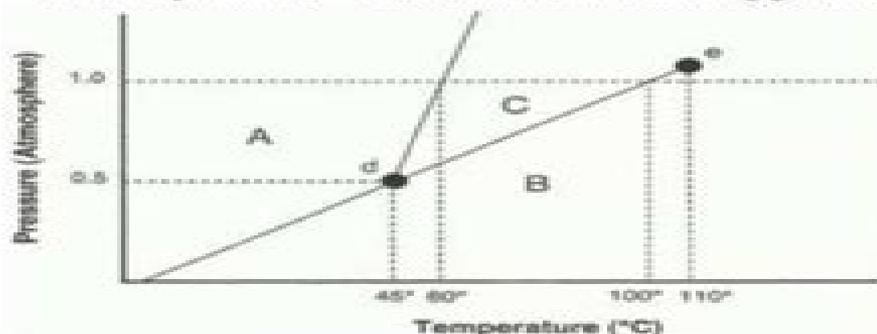


PHASE DIAGRAM REVIEW WORKSHEET

Part A – Generic Phase Diagram.

Answer questions 1 – 10 in relation to the following generic phase diagram.



1. Which section represents the solid phase? A
2. What section represents the liquid phase? C
3. What section represents the gas phase? B
4. What letter represents the triple point? a

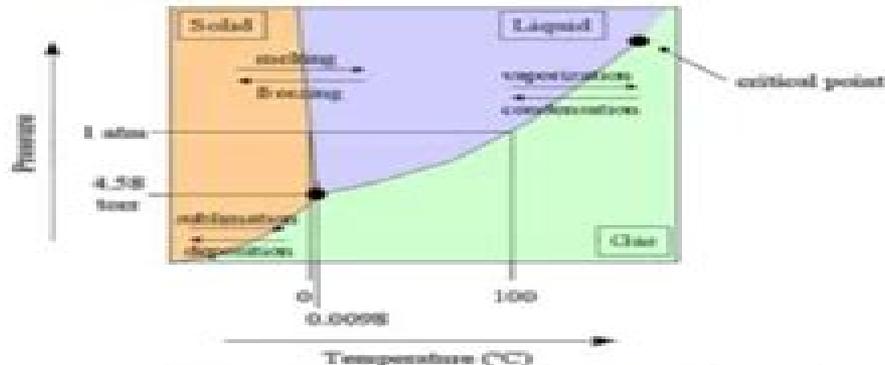
In your own words, what is the definition of a triple point?

answers will vary – s, l, g exist at equilibrium

5. What is this substance's normal melting point, at 1 atmosphere of pressure? 60°C
6. What is this substance's normal boiling point, at 1 atmosphere of pressure? 100°C
7. Above what temperature is it impossible to liquefy this substance, no matter what the pressure? 110°C
8. At what temperature and pressure do all three phases coexist? 0.5 atm, 45°C
9. At a constant temperature, what would you do to cause this substance to change from the liquid phase to the solid phase? increase the pressure
10. What does sublimation mean? to go from a solid to a gas directly

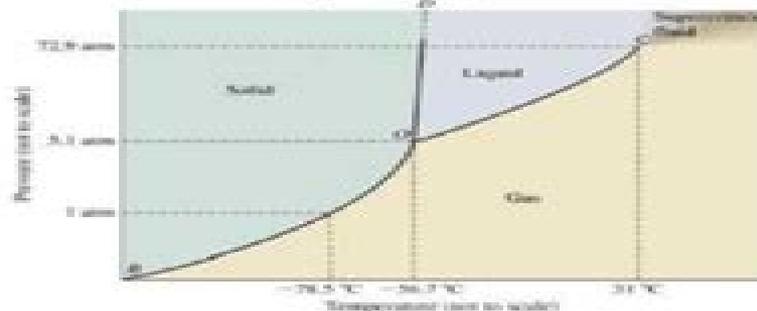
Part B – Phase Diagram for Water.

11. At a pressure of 1 atmosphere, what is the normal freezing point of water? 0°C
12. What is the normal boiling point of water, at one atmosphere of water? 100°C
13. People in Albuquerque live approximately 5,500 feet above sea level, which means the normal atmospheric pressure is less than 1 atm. In Albuquerque, will water freeze at a lower temperature or a higher temperature than at 1 atmosphere? higher Will water boil at a higher or lower temperature, than at 1 atmosphere? lower



Part C – Phase Diagram for Carbon Dioxide.

14. At 1 atmosphere and room temperature (25°C), would you expect solid carbon dioxide to melt to the liquid phase, or sublime to the gas phase? sublime to gas
15. Some industrial processes require carbon dioxide. The carbon dioxide is stored on-site in large tanks as liquid carbon dioxide. Assuming we lived at sea level (1 atm), how could carbon dioxide be liquefied? Increase the pressure



Multiple Choice Questions Phase Diagram

Richard H. Langley



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Table of Contents Multiple Choice Questions Phase Diagram

1. Understanding the eBook Multiple Choice Questions Phase Diagram
 - The Rise of Digital Reading Multiple Choice Questions Phase Diagram
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions Phase Diagram
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Questions Phase Diagram
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions Phase Diagram
 - Personalized Recommendations
 - Multiple Choice Questions Phase Diagram User Reviews and Ratings
 - Multiple Choice Questions Phase Diagram and Bestseller Lists
5. Accessing Multiple Choice Questions Phase Diagram Free and Paid eBooks
 - Multiple Choice Questions Phase Diagram Public Domain eBooks
 - Multiple Choice Questions Phase Diagram eBook Subscription Services
 - Multiple Choice Questions Phase Diagram Budget-Friendly Options

6. Navigating Multiple Choice Questions Phase Diagram eBook Formats
 - ePub, PDF, MOBI, and More
 - Multiple Choice Questions Phase Diagram Compatibility with Devices
 - Multiple Choice Questions Phase Diagram Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Multiple Choice Questions Phase Diagram
 - Highlighting and Note-Taking Multiple Choice Questions Phase Diagram
 - Interactive Elements Multiple Choice Questions Phase Diagram
8. Staying Engaged with Multiple Choice Questions Phase Diagram
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Choice Questions Phase Diagram
9. Balancing eBooks and Physical Books Multiple Choice Questions Phase Diagram
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Choice Questions Phase Diagram
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Questions Phase Diagram
 - Setting Reading Goals Multiple Choice Questions Phase Diagram
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Questions Phase Diagram
 - Fact-Checking eBook Content of Multiple Choice Questions Phase Diagram
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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