

**SAMPLE 1**

How many significant figures does 0.007 09 kg have?

All nonzero digits are significant. The zero between the 7 and 9 is significant. The zeros to the left of the decimal point are not significant. The quantity 0.007 09 kg has 3 significant figures.

**SAMPLE 2**

Divide 79.7 g by 0.89 cm<sup>3</sup>.

The quantity 79.7 g has 3 significant figures, but 0.89 cm<sup>3</sup> has only 2 significant figures. So, the product 8.955 056 18 g/cm<sup>3</sup> must be rounded to 2 significant figures. The rounded quantity is 9.0 g/cm<sup>3</sup>.

**SAMPLE PROBLEM A**

A sample of aluminum metal has a mass of 8.4 g. The volume of the sample is 3.1 cm<sup>3</sup>. Calculate the density of aluminum.

**SOLUTION**

Given: mass ( $m$ ) = 8.4 g  
 volume ( $V$ ) = 3.1 cm<sup>3</sup>  
 Unknown: density ( $D$ )

$$\text{density} = \frac{\text{mass}}{\text{volume}} = \frac{8.4 \text{ g}}{3.1 \text{ cm}^3} = 2.7 \text{ g/cm}^3$$

**SAMPLE PROBLEM B**

Express a mass of 5.712 grams in milligrams and in kilograms.

**SOLUTION**

Given: 5.712 g  
 Unknown: mass in mg and kg

The expression that relates grams to milligrams is

$$1 \text{ g} = 1000 \text{ mg}$$

The possible conversion factors that can be written from this expression are

$$\frac{1000 \text{ mg}}{\text{g}} \quad \text{and} \quad \frac{1 \text{ g}}{1000 \text{ mg}}$$

To derive an answer in mg, you'll need to multiply 5.712 g by 1000 mg/g.

$$5.712 \text{ g} \times \frac{1000 \text{ mg}}{\text{g}} = 5712 \text{ mg}$$

This answer makes sense because milligrams is a smaller unit than grams and, therefore, there should be more of them.

The kilogram problem is solved similarly.

$$1 \text{ kg} = 1000 \text{ g}$$

Conversion factors representing this expression are

$$\frac{1 \text{ kg}}{1000 \text{ g}} \quad \text{and} \quad \frac{1000 \text{ g}}{\text{kg}}$$

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