

2nd Year Engineering (COP) 19 and 20 class, Semester Fourth, 2021

Chem

**ACADEMIC
DEPT.**

Write in Short:

- 1) Define the term coordination number. **(2M)**
 Coordination number is the number of atoms or ions directly bonded to a central atom or ion in a complex.
- 2) Name any two coordination compounds. **(2M)**
 $[Co(NH_3)_6]Cl_3$ and $[PtCl_2(NH_3)_2]$
- 3) How would you distinguish between weak and strong field ligand complexes? **(2M)**
 By using spectrochemical series.
- 4) How are strong and weak field ligands? **(2M)**
 They are σ donor ligands.
- 5) What are the conditions for high spin and low spin complexes? **(2M)**
 d^4, d^5, d^6, d^7 complexes.
- 6) What is the difference between paramagnetic and diamagnetic? **(2M)**
 Paramagnetic has unpaired electrons, diamagnetic has paired electrons.
- 7) What is the term crystal field splitting? **(2M)**
 Energy level of d-orbitals in a complex is split into two groups: e_g and t_{2g} orbitals.
- 8) How would you measure the crystal energy of a complex? **(2M)**
 By using spectrochemical series.
- 9) What is the coordination number of $[Co(NH_3)_6]Cl_3$? **(2M)**
 6.
- 10) How would you measure the bond length of a complex? **(2M)**
 By using X-ray diffraction.
- 11) Discuss the crystal field splitting of d^4 complex with an example. Also, draw the orbital diagram of d^4 complex. **(4M)**
 In d^4 complex, the d -orbitals are split into two groups: e_g and t_{2g} . The t_{2g} orbitals are lower in energy than the e_g orbitals. The d^4 complex has four electrons in the d -orbitals. The first three electrons occupy the t_{2g} orbitals, and the fourth electron occupies the e_g orbitals. The orbital diagram is as follows:

Complex	Crystal field splitting	Orbital diagram
$[Co(NH_3)_6]Cl_3$ (high spin d^6)	$t_{2g}^4 e_g^2$	$t_{2g}^4 e_g^2$
$[PtCl_2(NH_3)_2]$ (low spin d^8)	$t_{2g}^6 e_g^2$	$t_{2g}^6 e_g^2$
$[Co(NH_3)_6]Cl_3$ (high spin d^6)	$t_{2g}^4 e_g^2$	$t_{2g}^4 e_g^2$

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