

Practice problems for Gravimetric analysis and gravimetric factor ratio Answers

1. Determine the gravimetric factor for the following sought analytes and weighed final products:

(a) Weighed final product: CaC_2O_4

Sought analyte: CaCO_3

$$\text{MW} (\text{CaC}_2\text{O}_4) = 40.08 \text{ g/mol} + 2(12.01 \text{ g/mol}) + 4(16.00 \text{ g/mol}) = 128.10 \text{ g/mol}$$

$$\text{MW} (\text{CaCO}_3) = 40.08 \text{ g/mol} + 12.01 \text{ g/mol} + 3(16.00 \text{ g/mol}) = 100.09 \text{ g/mol}$$

$$GF = \frac{1 \times \text{MW of CaCO}_3}{1 \times \text{MW of CaC}_2\text{O}_4} = \frac{100.09}{128.10} = 0.78134$$

(b) Weighed final product: PbI_2

Sought analyte: NaIO_3

$$\text{MW} (\text{PbI}_2) = 207.2 \text{ g/mol} + 2(126.90 \text{ g/mol}) = 461.0 \text{ g/mol}$$

$$\text{MW} (\text{NaIO}_3) = 22.99 \text{ g/mol} + 126.90 \text{ g/mol} + 3(16.00 \text{ g/mol}) = 197.89 \text{ g/mol}$$

$$GF = \frac{2 \times \text{MW of NaIO}_3}{1 \times \text{MW of PbI}_2} = \frac{2(197.89)}{461.0} = 0.8585$$

(c) Weighed final product: B_2O_3

Sought analyte: $\text{Na}_2\text{B}_4\text{O}_7$

$$\text{MW} (\text{B}_2\text{O}_3) = 2(10.81 \text{ g/mol}) + 3(16.00 \text{ g/mol}) = 69.62 \text{ g/mol}$$

$$\text{MW} (\text{Na}_2\text{B}_4\text{O}_7) = 2(22.99 \text{ g/mol}) + 4(10.81) \text{ g/mol} + 7(16.00 \text{ g/mol}) = 201.22 \text{ g/mol}$$

$$GF = \frac{1 \times \text{MW of Na}_2\text{B}_4\text{O}_7}{2 \times \text{MW of B}_2\text{O}_3} = \frac{201.22}{2(69.62)} = 1.445$$

(d) Weighed final product: Hg_2I_2

Sought analyte: HgO

$$\text{MW} (\text{Hg}_2\text{I}_2) = 2(200.59 \text{ g/mol}) + 2(126.90 \text{ g/mol}) = 654.98 \text{ g/mol}$$

$$\text{MW} (\text{HgO}) = 200.59 \text{ g/mol} + 16.00 \text{ g/mol} = 216.59 \text{ g/mol}$$

$$GF = \frac{2 \times \text{MW of HgO}}{1 \times \text{MW of Hg}_2\text{I}_2} = \frac{2(216.59)}{654.98} = 0.66136$$

2. The hydrogen sulphide (H_2S) in a 50.0 g sample of crude petroleum was removed by distillation and collected in a solution of CdCl_2 . The precipitated CdS was then filtered, washed, and ignited to CdSO_4 . Calculate the percentage of H_2S in the sample if 0.108 g of CdSO_4 was recovered.

$$\text{MW} (\text{H}_2\text{S}) = 2(1.01) \text{ g/mol} + 32.07 \text{ g/mol} = 34.09 \text{ g/mol}$$

$$\text{MW} (\text{CdSO}_4) = 112.41 \text{ g/mol} + 32.07 \text{ g/mol} + 4(16.00 \text{ g/mol}) = 208.48 \text{ g/mol}$$

$$GF = \frac{1 \times \text{MW of H}_2\text{S}}{1 \times \text{MW of CdSO}_4} = \frac{34.09}{208.48} = 0.1635$$

$$\% \text{H}_2\text{S} = [(0.108 \text{ g}) \times (0.1635)/50.0 \text{ g} \times 100] = 0.0353\%$$

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M Woodhall



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Table of Contents Gravimetric Analysis Problems With Answers

1. Understanding the eBook Gravimetric Analysis Problems With Answers
 - The Rise of Digital Reading Gravimetric Analysis Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gravimetric Analysis Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gravimetric Analysis Problems With Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Gravimetric Analysis Problems With Answers
 - Personalized Recommendations
 - Gravimetric Analysis Problems With Answers User Reviews and Ratings
 - Gravimetric Analysis Problems With Answers and Bestseller Lists
5. Accessing Gravimetric Analysis Problems With Answers Free and Paid eBooks
 - Gravimetric Analysis Problems With Answers Public Domain eBooks
 - Gravimetric Analysis Problems With Answers eBook Subscription Services
 - Gravimetric Analysis Problems With Answers Budget-Friendly Options
6. Navigating Gravimetric Analysis Problems With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gravimetric Analysis Problems With Answers Compatibility with Devices
 - Gravimetric Analysis Problems With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gravimetric Analysis Problems With Answers
 - Highlighting and Note-Taking Gravimetric Analysis Problems With Answers
 - Interactive Elements Gravimetric Analysis Problems With Answers
8. Staying Engaged with Gravimetric Analysis Problems With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gravimetric Analysis Problems With Answers
9. Balancing eBooks and Physical Books Gravimetric Analysis Problems With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gravimetric Analysis Problems With Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gravimetric Analysis Problems With Answers
 - Setting Reading Goals Gravimetric Analysis Problems With Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Gravimetric Analysis Problems With Answers
 - Fact-Checking eBook Content of Gravimetric Analysis Problems With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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