

## Calorimetry and Hess's Law Post Lab Questions

**CIRCLE YOUR FINAL ANSWER and SHOW ALL WORK where applicable.**

1. Calculate the heats of combustion for the following reactions from the standard enthalpies of formation listed in Appendix 3 in your textbook:



2. From the standard enthalpies of formation, calculate  $\Delta H^\circ_{\text{rxn}}$  for the reaction



For  $\text{C}_6\text{H}_{12}(\text{l})$ ,  $\Delta H_f^\circ = -151.9 \text{ kJ/mol}$ .

3. Determine the amount of heat (in kJ) given off when  $1.26 \times 10^4 \text{ g}$  of ammonia are produced according to the equation



$$\Delta H^\circ_{\text{rxn}} = -92.6 \text{ kJ/mol}$$

Assume that the reaction takes place under standard-state conditions at  $25^\circ\text{C}$ .

# Hess Law Post Lab Questions

**M Planty**

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Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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