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Boyle's Law

- The Volume of a fixed mass of gas varies indirectly with the pressure at constant temperature
- $PV = k$
- $P_1V_1 = P_2V_2$

Kinetic Theory and Boyle's Law

- Pressure of a gas is caused by the molecules of the gas striking the walls of the container
- If the gas is compressed to half the volume it had, twice as many molecules are present in any given volume
 - Twice as many impacts per second on the walls of the container
 - pressure of the gas will double

Ex 1: A balloon filled with Helium has a volume of 457 mL at standard atmospheric pressure. After the balloon is released, it reaches an altitude of 6.3 km where the pressure is only 65.5 kPa. What is the volume of the balloon at this altitude?

$$P_1V_1 = P_2V_2$$

$$P \downarrow \frac{1}{2} \quad V \uparrow \times 2$$

$$101.3 \text{ kPa} \times 457 \text{ mL} = 65.5 \text{ kPa} \times V_2$$

$$V_2 = 706.8 \text{ mL}$$

Ex 2: Under a pressure of 673 mm Hg, a confined gas has a volume of 335 mL. If the pressure is increased until the volume is 252 mL, what is the new pressure, assuming the temperature remains constant?

$$P_1V_1 = P_2V_2$$

$$673 \text{ mmHg} \times 335 \text{ mL} = P_2 \times 252 \text{ mL}$$

$$P_2 = 828.2 \text{ mmHg}$$

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