

Half Modern Chemistry Review

CHAPTER 5: THE PERIODIC TABLE

The following pages contain the best (but not all) of the information for this chapter in class.

Focus on this content, but make sure to review other topics, activities, handouts, worksheets, etc.

If you already take advanced-level SCIENCE class, your school is listed for state or PAAS after class.

**** Best items will be bold, underlined, another underline of this content. ****

COURSES:

- College-level (only possible by university/college-level course, SC 121)
- Community-college-level (must be general and general SC 121)
- 5.1: Describe the organization of the periodic table; identify metals, metalloids, nonmetals, gases, and liquids. IT 41 42
- 5.2: Describe Periodic Table as atomic/elemental about the properties (chemical and physical), conductivity, energy, and atomic type of the elements. IT 41 42 43
- 5.3: Identify common metals, metalloids, and nonmetals; identify the representative groups metals groups (groups 1, 2, 13, 14, 15, 16, 17, 18) IT 41 42
- 5.4: Compare periodic types of atomic size, ionization energy, and electronegativity. IT 41 42

5.1: PERIODS OF THE PERIODIC TABLE

• Introductory

- periodic table = the fact that states that the repeating chemical and physical properties of elements change periodically, with the atomic numbers of the elements
- periodic table = an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column.
- horizontal = a row of the table with same number of elements, atomic numbers, same outer shell electrons
- vertical = one of the columns of table with same number of elements that decrease from top atomic number (Z) through transition (atomic number 1-10) to the periodic table

• Chapter Highlights

- The periodic table states that the physical and chemical properties of the elements are periodic functions of their atomic numbers.
- The periodic table is an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column/group/family.
- The columns in the periodic table are referred to as groups.

5.2: ELECTRON CONFIGURATIONS AND THE PERIODIC TABLE

• Introductory

- shell model = one of the elements of Group 1 of the periodic table (lithium, sodium, potassium, rubidium, cesium, and francium)
- alkaline-earth metal = one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, barium, and radium)
- transition element = one of the metals that use one of the inner shell electrons in the outer shell to bond.
- metalloid element = an element in the block or block of the periodic table
- halogen = one of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine); halogens combine with most metals to form salts.

Modern Chemistry Holt Ch 5 Review Answers

Judd E. Hollander



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