

Molecular Geometry

How can molecular shapes be predicted using the VSEPR theory?

Why?

When you draw a Lewis structure for a molecule on paper, you are making a two-dimensional representation of the atoms. In reality however, molecules are not flat—they are three-dimensional. The true shape of a molecule is important because it determines many physical and chemical properties for the substance. In this activity you will learn how to predict molecular shapes.

Model 1 – Lewis Structures

Lewis Structures	H_2CO 3 electron domains (3 bonding, 0 nonbonding)	3-D Molecular Shape
<input type="checkbox"/> 1. H_2CO 	BeF_2 2 electron domains (2 bonding, 0 nonbonding)	
<input type="checkbox"/> 2. BeF_2 	CH_4 4 electron domains (4 bonding, 0 nonbonding)	
<input type="checkbox"/> 3. CH_4 	NH_3 4 electron domains (3 bonding, 1 nonbonding)	
<input type="checkbox"/> 4. NH_3 	H_2O 4 electron domains (2 bonding, 2 nonbonding)	
<input type="checkbox"/> 5. H_2O 	CO_2 2 electron domains (2 bonding, 0 nonbonding)	
<input type="checkbox"/> 6. CO_2  Lone pair = **		

Examine the drawings in Model 1.

1. What does a solid line between two element symbols represent in the drawing of the molecules?
2. What does a pair of dots represent in the drawing of the molecules?
3. In the English language, what does the word "domain" mean? (Your group must come to consensus on this question.)

Molecular Geometry Pogil Answer Key

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