

Gas Laws Study Guide

Write the Law, Formula, Symbols, Units, and a sentence describing the meaning of the symbols and units.

- A sample of gas is held at a constant volume and pressure.
- The pressure of a gas is held constant and the volume is changed.
- A sample of gas is held at a constant volume and temperature.
- A sample of gas is held at a constant pressure and temperature.
- A sample of gas is held at a constant volume and the number of moles is changed.
- A sample of gas is held at a constant pressure and the number of moles is changed.

Name _____ Date _____

Gas Laws Study Guide

Boyle's Temperature: $T = \dots$ K $T = \dots$ $^{\circ}C$

Standard Pressure: $P = \dots$ atm $P = \dots$ mmHg $P = \dots$ kPa $P = \dots$ bar $P = \dots$ Pa

Standard Volume: $V = \dots$ L $V = \dots$ m³

Match the Formula to the Variable

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Combined Gas Law

$$P_{total} = P_1 + P_2 + P_3 + P_4$$

Dalton's Law

$$PV = nRT$$

Ideal Gas Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Boyle's Law

$$PV = nRT$$

Charles's Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Gay-Lussac's Law

Complete the following statements by filling in the blanks & writing the correct units.

As Pressure increases Volume decreases (increases) given that _____ is constant meaning that they have a direct (inverse) relationship.

As Volume increases Temperature increases (decreases) given that _____ is constant meaning that they have a direct (inverse) relationship.

As Pressure increases Temperature increases (decreases) given that _____ is constant meaning that they have a direct (inverse) relationship.

As the number of moles of gas in a container increases the Volume increases (decreases), the Pressure increases (decreases), and the Temperature increases (decreases).

Gas Laws Answer Study Guide Chemistry

Lingsheng Yao



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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

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