



6. Which of the images shown above depicts a phase that has definite volume but not definite shape?

- A) The one in the middle  
 B) The one on the left  
 C) The one in the middle and the one on the right  
 D) The one on the right

7. Assuming all other parameters remain constant, what happens to the pressure of a gas when the volume of the gas is increased?

- A) increases  
 B) decreases  
 C) fluctuates sinusoidally  
 D) remains the same

8. You collect 550 mL of argon gas at 23.0°C. What volume will the gas occupy at 46.0 °C if the pressure remains constant?

- A) 270 mL  
 B) 532 mL  
 C) 999 mL  
 D) 1104 mL

9. What happens to the pressure of a gas inside a container if the temperature of the gas decreases?

- A) The pressure increases.  
 B) The pressure does not change.  
 C) The pressure decreases.  
 D) The pressure cannot be predicted.

10. A sealed metal can full of air is heated from 20 °C to 120 °C. The can was initially at 1.0 atmosphere of pressure. The final pressure will be

- A) 6.0 atm  
 B) 0.007 atm  
 C) 0.72 atm  
 D) 1.3 atm

11. The combined gas law relates which of the following?

- A) pressure and volume only  
 B) temperature and pressure only  
 C) volume and temperature only  
 D) temperature, pressure, and volume

12. If I have 25 mL of a gas at a pressure of 2.1 atm and a temperature of 300 K, what will

the pressure be if I increase the temperature to 400 K and compress the gas to a volume

of 10 mL?

- A) 1.4 atm  
 B) 8.6 atm  
 C) 0.008 atm  
 D) 7 atm  
 E) none of these

13. A rigid container contains 2.50 mol  $O_2$ , 0.90 mol  $N_2$ , and 1.00 mol  $CO_2$ . The total pressure of the system is 200 kPa. The partial pressure exerted by the  $N_2$  in the mixture is

- A) 25.0 kPa  
 B) 50.0 kPa  
 C) 100 kPa  
 D) 125 kPa  
 E) 150 kPa

# Gas Laws Test Answers

**Christian Drosten**



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## **Table of Contents Gas Laws Test Answers**

1. Understanding the eBook Gas Laws Test Answers
  - The Rise of Digital Reading Gas Laws Test Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Test Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Laws Test Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Test Answers
  - Personalized Recommendations
  - Gas Laws Test Answers User Reviews and Ratings
  - Gas Laws Test Answers and Bestseller Lists
5. Accessing Gas Laws Test Answers Free and Paid eBooks
  - Gas Laws Test Answers Public Domain eBooks
  - Gas Laws Test Answers eBook Subscription Services
  - Gas Laws Test Answers Budget-Friendly Options
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- ePub, PDF, MOBI, and More
  - Gas Laws Test Answers Compatibility with Devices
  - Gas Laws Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Gas Laws Test Answers
    - Highlighting and Note-Taking Gas Laws Test Answers
    - Interactive Elements Gas Laws Test Answers
  8. Staying Engaged with Gas Laws Test Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Gas Laws Test Answers
  9. Balancing eBooks and Physical Books Gas Laws Test Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Gas Laws Test Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Gas Laws Test Answers
    - Setting Reading Goals Gas Laws Test Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Gas Laws Test Answers
    - Fact-Checking eBook Content of Gas Laws Test Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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