

Experimental Conditions			
	Reactants	Products	
Before reaction	 2 molecules $N_2$	 3 molecules $H_2$	0 molecules $NH_3$
After reaction	 1 molecule $N_2$	0 molecules $H_2$	 2 molecules $NH_3$

- In this reaction, only the hydrogen is completely used up.
- $H_2$  is the **limiting reagent**, or the reactant that determines the amount of product that can be formed by a reaction.

# Guided Reading 12 3 Limiting Reagent And Percent Yield

**Jessica J Manson**



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